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The Modern Interpretation of the Wittig Reaction Mechanism

Peter A. Byrne, a and Declan G. Gilheany* a

Abstract: The mechanism of the Wittig reaction has long been a contentious issue in organic chemistry. Even now, more than 50 years after its announcement, its presentation in many modern undergraduate textbooks is either overly simplified or entirely inaccurate. In this review, we gather together the huge body of evidence that has been amassed to show that the Li salt-free Wittig reactions of non-stabilised, semi-stabilised and stabilised ylides all occur under kinetic control by a common mechanism in which oxaphosphetane (OPA) is the first-formed and only intermediate. The numerous recent significant additions to the subject – including computational studies and experimental material pertinent to both steps of the reaction (OPA formation and its decomposition) are discussed in detail, and the currently accepted explanations for the source of the stereoselectivity in Wittig reactions are given. We also present other mechanistic proposals that have been made during the history of the Wittig reaction, and show how they are unable to account for all of the experimental evidence that is now available. Details of certain experimental facts to do with Wittig reactions in the presence of Li cation are also included, although the precise mechanistic details of such reactions are yet to be established conclusively. We make the case that a clear distinction should henceforth be made between the unknown “Li-present” and the now well-established “Li salt-free” Wittig mechanisms.

1. The Wittig Reaction: Introduction, utility and recent developments

The Wittig reaction is perhaps the most commonly used method for the synthesis of alkenes. Several excellent reviews on the topic have previously been written that contain extensive detail on the Wittig reaction mechanism. More recent reviews do not focus on the details of the mechanism. The reaction (see Scheme 1) occurs between a carbonyl compound (aldehyde or ketone in general, 2) and a phosphonium ylide (1) to give alkene (3) with a phosphorus oxide (4) as the by-product. The ylide can be represented by structures 1a (fully ionic ylide form) or 1b (ylene form), which should not be taken as resonance forms. Rather they should be understood as illustrating the ionic character of the P-C bond with a contribution to the stabilisation of the carbanion by phosphonium. There has been much discussion in the literature over the nature of this stabilisation both in ylides, and in similar bonds in phosphorus stabilised carbanions, and phosphine oxides, and iminophosphoranes. There seems to be a developing consensus that the P=O and P=N bonds are composed of a polar covalent bond combined with an electrostatic interaction. However for the ylides the evidence suggests that there may be negative hyperconjugation of the lone pair on the ylide α-carbon into the σ orbitals of the bonds of phosphorus to its other substituents, albeit that the P-C bond is still heavily polarised towards carbon. For convenience, ylides will be represented in the ylene form (1b) throughout this review.

Scheme 1. The Wittig Reaction. X, Y and Z may each be alkyl, aryl or alkoxyl, and need not necessarily be the same. R2 may be alkyl, aryl, vinyl, or an electron withdrawing group (e.g. an ester). The carbonyl reactant (2) may be formaldehyde (R1a = R1b = H), an aldehyde (R1a = alkyl/aryl, R1b = H), or a ketone (R1a = alkyl/aryl, R1b = alkyl/aryl).

There are a number of features of the Wittig reaction that help to make its use so widespread:
- It is regioselective – the alkene is invariably formed from the ylide α-carbon to the carbonyl carbon.
- Reaction conditions are usually mild, even in comparison to the extremely useful regiospecific phosphorus based olefination, the Wadsworth-Emmons reaction.
- The starting materials are frequently easily obtainable. Ylides are easily accessible by in situ deprotonation of the parent phosphonium salt, obtained in turn by alkylation of a phosphine. Certain salts and even some stabilised ylides (see later for definition) are now commercially available.
- Ylides are tolerant of a number of other functional groups, so the reaction is suitable for synthesis of complex molecules.
- The stereoselectivity of the reaction can often be directed towards the desired alkene isomer by expedient choice of the nature of the reactants used (see later).
Despite the fact that Wittig & Geissler first reported on the reaction 60 years ago, it still remains the subject of many modern publications. Amongst the recent developments reported in the literature are:

- New methods for generation of the ylide,\textsuperscript{18,19,20} 
- Reactions showing atypical stereoselectivity,\textsuperscript{21,22,23,24,25,26,27} which can be induced e.g. by the use of phosphonium ylides with modified "spectator substituents" on phosphorus,\textsuperscript{21,22,24,26} by the use of trialkylsililium base to generate the ylide,\textsuperscript{27} and by the addition of methanol at low temperature to the oxaphosphetane adducts from reactions of non-stabilised ylides.\textsuperscript{22} 
- New reaction conditions.\textsuperscript{25,28,29} 
- One-pot reaction from either phosphine,\textsuperscript{19,20,29,30,31} or alcohol\textsuperscript{32} 
- Synthetic studies.\textsuperscript{33} 
- Catalytic variants.\textsuperscript{34} 
- Our new chromatography-free method for the removal of the phosphine oxide by-product, yielding pure alkene product, and with optional conversion of the oxide to phosphine.\textsuperscript{35} 
- Reaction kinetics and relative nucleophilicity of ylides.\textsuperscript{36} 
- Mechanistic studies on reactions of ketones.\textsuperscript{37} 
- Demonstration by us of a unified mechanism for the Li salt-free Wittig reaction of all phosphonium ylide types from reactions of β-heteroatom-substituted aldehydes.\textsuperscript{38,39} 
- A study of the decomposition of the OPA intermediate (see later), which indicates that its decomposition may occur by a single step.\textsuperscript{40} 
- Computational studies on the reaction mechanism\textsuperscript{41,42,43} and on the stability of heteroatom-stabilised ylides including phosphonium ylides.\textsuperscript{44} 
- Recent synthetic applications including: photochromic dithienylethenes,\textsuperscript{45} enantioselective synthesis of highly substituted cycloalkanes by organocatalysed domino Michael-Wittig reactions,\textsuperscript{46} one-step synthesis of Z-allylic esters and alcohols by a SCOOPY-modified Wittig reaction,\textsuperscript{47,48} and stereoselective total synthesis of (-)-spirofungin A (involving several Wittig reactions).\textsuperscript{49} 

The scope of the related reaction of phosphonium ylides with N-sulfonyl imines has also been very effectively broadened, with the development of reactions of non-stabilised\textsuperscript{50} semi-stabilised\textsuperscript{51} and stabilised\textsuperscript{52} ylides that can be tuned to give complete E or Z selectivity by choice of the appropriate N-sulfonyl alkyl group.

2. Classification and selectivity trends of phosphonium ylides in the Wittig reaction

Phosphonium ylides are broadly categorised according to the nature of the substituent(s) attached to the α-carbon (R\textsuperscript{2} in Fig. 1). If R\textsuperscript{2} is an alkyl group, the ylide is referred to as being non-stabilised or reactive, as ylides of this type readily react with ambient moisture, and as such are not stable in air. If R\textsuperscript{2} is a phenyl or alkenyl group, the ylide is somewhat less prone to hydrolysis owing to the conjugative stabilisation by the unsaturated group, and so these ylides are called semi-stabilised ylides. Triphenylphosphine-derived ylides (X, Y, Z = Ph in Fig. 1) in which R\textsuperscript{2} is a carbonyl, ester, nitrile, sulfone or other such conjugating group are typically stable to hydrolysis in air. As a result, these ylides are referred to as stabilised ylides. This name has also come to refer to all ylides in which R\textsuperscript{2} is a conjugating group, even though certain ylides derived from phosphines other than triphenylphosphine are somewhat sensitive to hydrolysis in air.\textsuperscript{53}

Most ylides undergo Wittig reactions with carbonyl compounds very rapidly,\textsuperscript{5} although quantitative rate data on this subject is sparse. Where such data is available,\textsuperscript{36,53,54,55,56,57,58} quantitative comparison of ylide reactivities is difficult, since the Wittig reactions in different studies were often carried out under different conditions and, most significantly, with different solvents. However, qualitatively, it is known that non-stabilised ylides react essentially instantaneously (giving stable oxaphosphetane -- abbreviated OPA - at low temperature), while reactions of semi-stabilised ylides are generally complete in a matter of seconds or minutes (especially with aldehydes), even at low temperature. Some triphenylphosphine-derived stabilised ylides do not undergo Wittig reactions or require heating to effect the reaction,\textsuperscript{5} but we have observed methylidiphenylphosphine-driven stabilised ylides to react rapidly with aldehydes even at -78 °C.\textsuperscript{18,39} This is consistent with the observation of Frøyen that the reaction rates of fluorenylides with p-nitrobenzaldehyde increases with successive replacement of the P-phenyl groups of the ylide phosphonium moiety with P-ethyl groups.\textsuperscript{55}

A very significant aspect of the Wittig reaction is that, broadly speaking, the nature of the ylide used in a Wittig reaction dictates the stereoselectivity of the reaction. Certainly, in the case of triphenylphosphine-derived ylides it is true that in general the E or Z selectivity can be predicted based on the degree of anion stabilisation conferred on the ylide by its α-substituent (R\textsuperscript{2}). These trends are set out below. For reactions of ylides derived from other phosphines, prediction of the stereochemical outcome is not so straightforward and may be further complicated by the presence or absence of dissolved salts (especially Li salts, \textit{vide infra}).

Alkylidenetriphenylphosphoranes (Chart 1) are the most frequently used non-stabilised ylides. These show very consistently high Z-selectivity in Li-salt free reactions with a broad range of aldehydes.\textsuperscript{4} The only known exceptions involve reactions of ethylidenetriphenylphosphorane with aromatic aldehydes (and these reactions have been shown not to proceed under kinetic control - see below). Selectivity is generally highest in reactions with tertiary aldehydes, although in most cases even reactions of primary aldehydes still show overwhelming Z-selectivity for the alkene. By comparison, (alkylidene)alkylidiphenylphosphoranes (Chart 1 (b)) typically show very significantly reduced Z-selectivity; some reactions with primary aldehydes even showing moderate E-selectivity.\textsuperscript{5} Alkylidene-t-butyldiphenylphosphoranes (Chart 1 (c)) show comparable Z-selectivity to alkylidetriphenylphosphoranes.

![Fig. 1. Classification of ylides.](image-url)
Summary of observed selectivity trends for reactions of representative ylides. EWG = electron-withdrawing group (e.g. C(O)R, SO2R, CN etc.); R, R1, R2 = alkyl; X = aryl or aryl.

5 (Alkylidene)alkyl-dibenzo phospholanes (Chart 1 (d)) show extremely high E-selectivity. A generally applicable observation on the reactions of all of these non-stabilised ylides is that selectivity for Z-alkene is highest for a particular ylide in reactions with tertiary aldehydes, while reactions with primary aldehydes show the highest E-selectivity.

Wittig reactions of the most commonly encountered semi-stabilised stabilised ylides, derived from triphenyolphosphine, are generally not very selective for one isomer or another. Thus allylidene or benzylidene triphenyolphosphoranes (Chart 1 (e)) generally react to give approximately equal proportions of allylidene or benzylidene triphenylphosphoranes (Chart 1 (e)) generally not very selective for one isomer or another. Thus

Scheme 2. The betaine mechanism for the Wittig reaction.

10 Johnson's review. Par

3.1. The Betaine mechanism

55 Essential features: Nucleophilic addition of ylide to carbonyl to form betaine (diol species 5 in Scheme 2; similar to an aldol reaction), followed by ring closure to OPA (species 6) and decomposition of OPA to alkene and phosphine oxide (see Scheme 2).

60 At the time of the development of this mechanism by Wittig and co-workers, there was a significant amount of experimental evidence strongly suggesting the involvement of betaine intermediates in the Wittig reaction. There was also an absence of evidence to the contrary and the alternative explanations for the observed phenomena that have since been amassed had not yet been recognised. The experimental observations informing the betaine rationale, the alternative explanations for these observations, and the evidence that indicates the non-involvement of betaines is summarised below.

(i) Observation indicating the involvement of betaine: The most significant evidence came from the Wittig reactions of several ylides (generated using phenyllithium from the parent phosphonium bromide salt) with carbonyl compounds at low temperature, which gave a precipitate that on warming yielded alkene and phosphine oxide but if treated with HBr gave β-hydroxyphosphonium salt (β-HPS, an example of which is species 7 in Scheme 3). Subsequently the precipitate was isolated and shown to be a betaine-LiBr complex (which has since been detected spectroscopically).

Scheme 2).

65 Alternative explanation of observation: Since the above experiments were carried out, it has been conclusively shown that solutions containing only OPA (as confirmed by NMR) and react with LiBr to give a betaine complex 62 and 63 of this mechanism.

60 Johnson's review.
Rebuttal / alternative explanation: The epoxide cleavage and β-HPS deprotonation experiments necessarily produce betaine and, plainly, betaine gives OPA. However, while a betaine-LiBr complex has been observed by NMR, uncomplexed betaines have never been detected spectroscopically in a Wittig reaction mixture, or in an independent betaine generation experiment. OPAs are also necessarily produced in each of the independent generations of betaine, and they can be detected by NMR in solution in such experiments, and in Wittig reactions. Therefore experiments that necessarily involve betaine and that ultimately give alkene and phosphine oxide do not mean that they must be involved en route from ylide and aldehyde to alkene and phosphine oxide in the Wittig reaction. To be excluded as a Wittig intermediate, it is sufficient if betaine lies at a relatively higher energy than OPA and that there be a plausible alternative mechanistic route to OPA (which there is – mechanism 8).

(iii) Explanation of stereoselectivity of reactions of non-stabilised ylides within the context of the Wittig betaine mechanism: The universally high Z-selectivity in Li salt-free Wittig reactions of aldehydes with alkyldienetriphenylphosphoranes was rationalised as being a consequence of addition of the ylide to the carbonyl in such a way as to give anti-betaine (with a P-C-C-O dihedral angle of 180°, as shown in Fig. 2). The minimisation of steric repulsion in the transition state (TS) leading to an anti-betaine would dictate that there should also be a 180° dihedral angle between the vicinal R1 and R2 groups, thus giving an anti-erythro-betaine, which could undergo bond rotation to the syn-conformation and ring-close to cis-OPA, which ultimately leads to Z-alkene. So the high Z-selectivity was proposed to be due to kinetically favoured formation of erythro-betaine. Consistent with this interpretation was that deprotonation of erythro-(1-hydroxy-1-phenylprop-2-yl)triphenylphosphonium bromide with BuOK in diethylether gave the Z-alkene product almost exclusively, indicating high stereospecificity (and thus irreversibility) of conversion of the erythro-betaine, and cis-OPA necessarily produced, to Z-alkene under Li-salt-free conditions (see Scheme 3). Similar experiments in the presence of Li showed considerable conversion to the E-alkene (Scheme 3).

Rebuttal and alternative explanation: Another explanation ([2 + 2] cycloaddition) exists that is consistent with all of the observations that support the betaine mechanism, but which also accounts for several other facts that the betaine mechanism cannot explain. This is discussed in sections 3.8 and 4. However, it is appropriate to briefly state at this point that according to the betaine rationale, all Wittig reactions of non-stabilised ylides – even those that are not derived from triphenylphosphine – should be irreversible and highly Z-selective. In fact, Wittig reactions of non-stabilised ylides derived from other phosphines frequently show little stereoselectivity, and in some cases very high E-selectivity has even been observed. Irreversibility of OPA formation has been demonstrated for selected examples of these reactions (ones showing low selectivity). Thus, the betaine mechanism is unable to account for moderate to low Z-selectivity under kinetic control in Wittig reactions of non-stabilised ylides. Furthermore, the reactions of certain non-stabilised ylides (primarily ethylides – see later) are the only ones that are known not to be irreversible under Li salt-free conditions. It is also pertinent that OPA is the only observable intermediate in reactions of non-stabilised ylides, and that betaines have been shown conclusively not to be involved in the Wittig reactions of dibenzophosphole-derived non-stabilised ylides.

(iv) Explanation of stereoselectivity of reactions of stabilised ylides within the context of the Wittig betaine mechanism: Wittig reactions of stabilised ylides derived from triphenylphosphine are usually highly E-selective. The reaction of triphenylphosphine with ethyl-trans-2-phenylglycidate in refluxing ethanol in the presence of m-chlorobenzaldehyde gave a large amount of m-chlorocinnamate (crossover product) as well as the expected ethyl Z-cinnamate. This was interpreted to be occurring by the mechanism shown in Scheme 4. Subjecting the cis-epoxide to the same reaction conditions gave E-cinnamate and a smaller proportion of crossover product.

These observations were rationalised by postulating an increased propensity for equilibration of the kinetically favoured erythro-betaine and/or cis-OPA in reactions of these ylides, so that there would be equilibration of the intermediates by reversion to Wittig reactants. The increased reversion was supposed to be as a result of the greater stability, and thus longer lifetime, of the intermediates in reactions of stabilised ylides. If this scenario were true, greater E-selectivity should be observed if the irreversible decomposition to alkene and phosphine oxide.

Scheme 3. (a) Experiments of Schlosser and Christmann on deprotonation of β-HPSs derived from non-stabilised ylide.64 65 (b) Fischer projections of Schlosser’s β-HPSs, showing that the β-HPS derived from cis-OPA is erythro, and that derived from trans-OPA is threo.

Fig. 2. Two diagrams of the transition state leading to the anti-conformer of erythro-betaine, which has been proposed as the kinetically favoured path for the Wittig reaction.

in keeping with the observation of lower Z-selectivity in Wittig reactions conducted in the presence of Li.

cleavage of epoxide (of defined stereochemistry) with phosphine,66 or by deprotonation of β-HPS (of defined stereochemistry),54,62,64,67 and were shown to decompose to alkene and phosphine oxide (Scheme 3(a)).
were faster for trans-OPA than for cis-OPA, or if OPA formation were irreversible and ring-closure was faster for the threo-betaine than for the erythro isomer.

**Rebuttal and counter-evidence:** The betaine mechanism requires reversible formation of betaine and/or OPA in reactions of stabilised ylides in order to explain the observed selectivity – it is unable to account for E-selectivity under kinetic control. It has been shown that deprotonation of erythro-β-HPs derived from ester-stabilised ylides decompose stereospecifically to Z-alkene (presumably this necessarily occurs through erythro-betaine and then cis-OPA). Since neither the betaine nor the OPA precursor of Z-alkene underwent reversion, it follows that the highly E-selective reactions of these ester-stabilised ylides are under kinetic control. Thus the betaine mechanism cannot be in operation in these reactions. See section 4 for full details. The positive crossover results in the epoxide-cleavage experiments of Scheme 4 were obtained under conditions in which betaine was necessarily generated (and at a temperature much higher than is typically used for Wittig reactions), and do not mean that betaine is an essential intermediate in Wittig reactions of stabilised ylides, especially since it has been demonstrated that both high E and high Z-selectivity are possible under kinetic control in reactions of these ylides.

**Explanation of stereoselectivity of reactions of semi-stabilised ylides within context of the betaine mechanism:** Reactions of triphenylphosphine-derived semi-stabilised ylides show little or no selectivity (i.e. significantly decreased Z-selectivity relative to reactions of non-stabilised ylides). Deprotonation of erythro-(2-hydroxy-1,2-diphenylethyl)(methyl)diphenylphosphonium by "BuLi (Li present) in THF showed almost completely stereospecific conversion to Z-alkene, indicating that the corresponding Wittig reaction is likely to be under kinetic control (see Scheme 5). However, deprotonation experiments using the same β-HPs in ethanol with sodium ethoxide base in the presence of reactive aldehyde m-chlorobenzaldehyde gave a large amount of crossover product, while the stilbene that did form showed a Z/E ratio of around 90:10, indicating significant reversal of either erythro-betaine or cis-OPA to ylide and aldehyde under the reaction conditions (see Scheme 5).

These results were rationalised in a similar fashion to reactions of stabilised ylides; the increased stability of the ylide (vs. non-stabilised) was supposed to translate to the betaine and/or OPA intermediate, which was thought to lead to increased reversion of one of the intermediates to ylide and aldehyde. The reversion to starting materials was thought to be less prevalent in stabilised ylides since the intermediates derived from the latter should be more stable and thus longer-lived. As with stabilised ylides, reversion of the intermediate(s) would lead to greater E-selectivity if the decomposition to products were faster for trans-OPA than for cis-OPA, or if ring-closure of the betaine were irreversible, and faster for the threo-betaine than for the erythro isomer.

**Rebuttal and counter-evidence:** The same arguments that were presented in point (iv) above (stabilised ylides) apply here. Stereoselective decomposition of both erythro and threo-β-HPs (via betaine and OPA – the latter of which was observed by NMR) indicates that reactions of semi-stabilised ylides are irreversible. Relevant to the positive crossover experiment of Scheme 4 is a report that, at least in certain circumstances, the addition of methanol at low temperature (i.e. before OPA decomposition occurs) to the Wittig reactions of non-stabilised ylides causes very high E-selectivity in the reactions, which in the absence of methanol (or if it is added after warming to room temperature) would show high Z-selectivity.

**3.2. Bergelson’s “C-P-O-C” betaine mechanism**

**Essential features:** Nucleophilic attack of carbonyl oxygen at phosphorus of ylide to give a “C-P-O-C” betaine (with charges on the carbon atoms) followed by ring closure to OPA and decomposition of OPA to alkene, as shown in Scheme 6. A similar proposal was made by Schneider, which took account of the trigonal bipyramidal geometry at phosphorus in the proposed transition state (TS), who postulated that the observed Z-selectivity in reactions of non-stabilised ylides was a consequence of the steric effects that would result in such a TS.

**Evidence against this mechanism:** Rate-determining attack of the carbonyl at phosphorus should show a negative ρ value for carbonyl reactants since a positive charge is developed at the carbonyl carbon. However, carbonyl compounds have been experimentally found to show positive ρ for the Wittig reaction in reactions of all ylide types. Furthermore, no dipolar intermediate can be observed by NMR in Wittig reactions.
Scheme 6. Alternative betaine mechanism.

3.3. Schweizer mechanism

Essential features: On the basis of the observation that certain semi-stabilised and stabilised phosphonium ylides and conjugated carbonyl compounds reacted in alcohol solvents to give vinylphosphine oxides in addition to the expected alkene and phosphine oxide products, Schweizer et al. proposed that in this medium, these reactants initially form a betaine, which becomes protonated by the alcohol, and then undergo net elimination of water to give a vinylphosphonium salt. This could undergo nucleophilic attack at phosphorus by either alkoxy or hydroxide. The resultant alkoxy or hydroxy phosphonium salt could expel either the vinyl group as alkene, giving triphenyolphosphine oxide (the expected Wittig product), or benzene, giving vinyl phosphine oxide. See Scheme 7.

Evidence against this mechanism: Nucleophilic attack at phosphorus should result in at least partial inversion of the configuration at phosphorus. It has been shown that Wittig reactions of (benzylidene)ethylmethylphenylphosphorane with benzaldehyde give phosphine oxide with retention of configuration at phosphorus in ether solvent, while the same reaction (in which the ylide was generated from the parent phosphonium salt using ethoxide base) in ethanol was shown subsequent to the publication of Schweizer’s paper to proceed with retention of configuration at phosphorus. It was thus proposed that formation of Wittig products by OPA formation (from betaine) and the formation of vinylphosphine oxide occur by separate pathways.

It has since been shown that, at least in certain circumstances, the addition of methanol at low temperature (i.e. before OPA decomposition occurs) to the Wittig reactions of non-stabilised ylides causes very high E-selectivity in the reactions, which in the absence of methanol (or if it is added after warming to room temperature) show high Z-selectivity. This was proposed to be as a result of β-HPS formation from OPA and methanol, and deprotonation of the β-HPS by methoxide to give β-hydroxy ylide (10 in Scheme 8), which can re-form OPA (non-stereospecifically) by proton transfer from the hydroxyl group to the ylidy carbon (see Scheme 8). It is likely that the vinylphosphonium salts produced in the study of Schweizer et al. were produced from reaction of OPA, not betaine, with ethanol.

3.4. Olah’s single electron transfer mechanism

Essential features: Attempted Wittig reactions in refluxing solvents of non-stabilised ylides with adamantane-2-one, hydroxyadamantan-2-one, bicyclo[3,3,1]nonan-9-one or benzophenone gave the carbonyl reduction product (alcohol) and phosphonium salt as the only products after work-up, or in addition to the Wittig products. Based on these observations, the Wittig reaction was proposed to proceed by a one electron transfer mechanism, i.e. the transfer of an electron from the ylide to the carbonyl compound to initially give a tight radical ion pair, supposed to be in equilibrium with a P-O bonded diradical species (see Scheme 9), which would subsequently form betaine and then alkene. The alcohol formation was particularly favoured in the reactions of sterically bulky ylides isopropylidenetriphenylphosphorane and diphenylmethylenetriphenylphosphorane. The hydrogen source was the reaction solvent, as reactions in toluene gave benzylated toluenes. The one electron transfer mechanism was also advocated by Yamataka and co-workers on the basis of their observation that there is no significant kinetic isotope effect for the reaction of isopropylidenetriphenylphosphorane with benzaldehyde having a 14C-labelled carbonyl group.

Evidence against this mechanism: A number of tests for radical involvement in Wittig reactions are described in the review of Vedejs and Peterson. One of these is reproduced here. The unimolecular rate constants for ring opening of cyclopropylcarbonyl radicals are extremely large (Scheme 10(b)). By analogy, if the formation of radical ions from the Wittig reactants shown in Scheme 10(a) was occurring (see path B) then exceptionally fast ring cleavage would occur. Instead, the reaction gives the expected Wittig product in high yield with the expected high Z-selectivity (see Scheme 10(a), path A). This strongly implies that electron transfer is not an intrinsic part of the Wittig reaction mechanism, although electron transfer may be possible between suitable Wittig reactants under the right experimental conditions. In cases where it is possible, this probably leads to the formation of side-products. In one of their publications, Vedejs and Marth list a series of examples of
Scheme 9. Proposed mechanism for the Wittig mechanism involving single electron transfer from ylide to carbonyl species.

Scheme 10. (a) Test for involvement of radical species in the Wittig reaction, (b) Rates of ring opening for cyclopropylcarbonyl radicals.

5 Essential features: Direct cis-selective cycloaddition of ylide and carbonyl to give OPA with oxygen in an axial position in the phosphorus-centred trigonal bipyramid (although no rationale for the cis-selectivity was presented), followed by pseudorotation about phosphorus to place oxygen in an equatorial site, subsequent cleavage of the P-C bond to give a "P-O-C-C" betaine and finally scission of the O-C bond to give alkene and phosphine oxide, as shown in Scheme 11. Betaines derived from non-stabilised ylides were supposed to be very short-lived and thus to quickly decompose to Wittig products, whereas betaines with a stabilising group on the carbanion carbon (i.e. derived from a stabilised ylide) were thought to be longer lived due to the greater stabilisation of the negative charge, so that rotation could occur about the C-C bond before C-O bond scission, thus giving rise to E-alkene.

3.5. Bestmann’s “P-O-C-C” betaine mechanism

Essential features: Direct cis-selective cycloaddition of ylide and carbonyl to give OPA with oxygen in an axial position in the phosphorus-centred trigonal bipyramid (although no rationale for the cis-selectivity was presented), followed by pseudorotation about phosphorus to place oxygen in an equatorial site, subsequent cleavage of the P-C bond to give a "P-O-C-C" betaine and finally scission of the O-C bond to give alkene and phosphine oxide, as shown in Scheme 11. Betaines derived from non-stabilised ylides were supposed to be very short-lived and thus to quickly decompose to Wittig products, whereas betaines with a stabilising group on the carbanion carbon (i.e. derived from a stabilised ylide) were thought to be longer lived due to the greater stabilisation of the negative charge, so that rotation could occur about the C-C bond before C-O bond scission, thus giving rise to E-alkene.

Evidence against this mechanism: OPA formation has been shown to be irreversible and stereospecific for Wittig reactions of all ylide types in all but a small and well-defined set of reactions (of non-stabilised ylides; see section 4.1 part (i)) by stereospecific decomposition of β-HPS to alkene. Thus there can be no equilibration of intermediates after OPA has been formed (except in the cases detailed in section 4.1 part (ii)), and intermediates such as the betaine shown in Scheme 12 can play no part in the Wittig reaction.

6 Essential features: The first mechanism involved an entity with a C-C bond and an unpaired electron on each of phosphorus and oxygen, presumably formed by transfer of one electron from each of the C=O and P=C bonds into the new C-C bond, and movement of the remaining electron from each bond on to oxygen and phosphorus, respectively. The second involved a P-O bond with two carbon-centred radicaloids formed by an analogous process (this differs from the mechanism of Ohal, in which electron transfer from ylide to carbonyl was suggested to occur before the occurrence of P-O bonding). The kinetically favoured pathway was postulated to favour cis-OPA formation due to orthogonal approach of the ylide and aldehyde in the diradical intermediate, which was proposed for non-stabilised ylides to be short lived, while it would be longer lived for stabilised ylides and could thus undergo bond rotation to give trans-OPA and E-alkene.

Evidence against this mechanism: The first case, involving the C-C bonded diradical, can be ruled out by the same experiment that quashed the involvement of betaines in reactions of non-stabilised ylides (mentioned in section 3.1 part (ii), and discussed fully in section 4.2 below). The second, involving the P-O bonded diradical, can be ruled out for the same reasons as
were invoked for the Bergelson-Schneider mechanism (see section 3.2 above), and in particular by the fact that this mechanism would predict a negative $\rho$ value for carbonyl compounds in the Wittig reaction where a positive value is observed. Furthermore, no evidence has been reported in support of the existence of either of the proposed diradical intermediates.

### 3.7. Schlosser “leeward approach” mechanism

Schlosser and Schaub proposed the formation of OPA by [2 + 2] cycloaddition of ylide and aldehyde, with the mechanistic details differing somewhat from its original proposal by Vedejs and Snoble, and from later variants by Vedejs & co-workers (see section 3.8).

**Essential features:** The kinetically favoured OPA-forming TS was postulated to be planar, with trigonal bipyramidal geometry about phosphorus (i.e. complete reorganisation had occurred of the substituents about phosphorus to the new geometry). The ylide $\alpha$-substituent was supposed to influence the arrangement of the $P$-phenyl groups (1-2 interactions) in such a way as to make steric interactions of the carbonyl substituent with the phosphorus substituents (1-3 interactions) much greater in the TS leading to trans-OPA (see Fig. 3(a)) than in the cis-selective TS (Fig. 3(b)).

This would mean that the cis-OPA would be favoured kinetically - hence its preferential formation in Wittig reaction of non-stabilised ylides.

**Evidence against this mechanism:** Based on the above rationale for the cycloaddition TS, it follows that cis-OPA should be the most thermodynamically stable OPA isomer. However, there exists no known example of isomerisation of trans-OPA to cis-OPA (whether by reversal to Wittig reactants or otherwise). On this basis it can be concluded that trans-OPA is thermodynamically favoured over cis-OPA, which is intuitively sensible, as the former should have similar 1-3 steric interactions but significantly decreased 1-2 interactions compared to the latter. Therefore the cycloaddition TS cannot be product-like (i.e. it must be early), as otherwise it would be trans-selective, and so could not have a trigonal bipyramidal arrangement of the substituents about phosphorus - an arrangement that is intrinsic to the explanation of selectivity in the Schlosser mechanism. Also, it is difficult to rationalise $E$-selectivity in reactions of stabilised ylides in the context of this mechanism, involving as it does direct OPA formation, given that OPA formation has been shown to be irreversible for reactions of all ylide types. Although the rationale put forth to explain the observed selectivity in the context of this mechanism is not consistent with all experimental observations, it is close to the currently accepted mechanism in that it involves OPA formation by direct [2+2] cycloaddition of ylide and aldehyde.

### 3.8. Vedejs cycloaddition mechanism

Vedejs and Snoble were the first to advance the proposal of direct irreversible cycloaddition of the ylide and aldehyde to give OPA, followed by irreversible and stereospecific cycloreversion of the OPA to give phosphine oxide and alkene (see Scheme 13). Later, having adduced the evidence that established that all Li salt-free Wittig reactions are irreversible (see section 4.1), this [2 + 2] cycloaddition mechanism was elaborated by Vedejs and co-workers to include details of how the stereochemistry of the product alkene is decided at the TS of the cycloaddition step (see Fig. 4). Indeed, it is only recently that definitive computational and experimental results have appeared that shed light on the elusive details of this part of the mechanism (see the section on the modern mechanistic interpretation for the Wittig reaction). In any case, it is the origin of the stereochemistry induced in the OPA intermediate that is of greatest interest since this stereochemistry is retained exactly in the alkene product after decomposition, and this was the primary focus in the work of Vedejs and co-workers.

Reactions of aldehydes with non-stabilised alkyldienetriphenylphosphoranes were suggested to proceed through an early puckered TS (meaning that the carbonyl C=O bond and ylide P=C bond approach each other at a relatively large angle) in which both bond formation and rehybridisation about the atoms in the forming ring are at an early stage (see Fig. 4(a)). The fact that the four substituents bound to phosphorus in the ylide remain in a nearly tetrahedral arrangement about phosphorus in the TS has particular consequences for the shape of the TS. This disposition of the phosphorus substituents means that planar approach of the ylide and aldehyde is disfavoured for two reasons – firstly one of the $P$-phenyl groups must necessarily project in the direction of the forming P-O bond, and secondly there exists the possibility of large 1-3 steric interactions between the phosphorus substituents and the aldehyde substituent. Puckering relieves both of these unfavourable interactions, and minimises 1-2 steric interactions (between aldehyde and ylide substituents) if the large substituent on the carbonyl is placed in a pseudo-equatorial position and the ylide $\alpha$-substituent is pseudo-axial in the forming ring. This energetically favoured TS leads to cis-OPA. The lowest energy trans-selective TS would necessarily be significantly higher in energy whether puckered or planar. Thus, cis-OPA is thought to be formed selectively under conditions of kinetic control, and to decompose irreversibly and stereospecifically to Z-alkene and phosphine oxide by syn-cycloreversion.

Reactions of stabilised ylides were proposed by Vedejs and co-workers to proceed through a later transition state in which
Fig. 4. (a) Proposed cis-selective cycloaddition transition state in reactions of alkylidenediarylphosphoranes (b) Proposed trans-selective cycloaddition transition state in reactions of stabilised ylides in Vedejs mechanism.  


The finer details of the mechanism, including unresolved issues related to the second step of the mechanism and especially the experimental and computational evidence that has been amassed in its favour, are described fully in the next section.

4. Modern mechanistic interpretation of the Wittig reaction

There are many questions to be answered in considering how the reaction of a phosphonium ylide and aldehyde progresses to give an alkene and phosphine oxide progresses. These include:

(i) Is the reaction under kinetic or thermodynamic control?

(ii) What is the nature of the first-formed intermediate? Many have been proposed, but OPAs and betaines are the only ones for which sufficient experimental evidence has been presented to enable them to gain widespread acceptance.

(iii) What is the origin of the observed pattern of Z/E-selectivity? This is intrinsically linked to the first two points.

(iv) Much more recently, the process by which the OPA intermediate decomposes to alkene & phosphine oxide has come into question. Does this process involve two steps – pseudorotation to a “C-apical” intermediate followed by cycloreversion – or just one step (cycloreversion), without formation of intermediate(s)?

4.1. The operation of kinetic control in the Wittig reaction

The substantial evidence available on this issue strongly suggests that, in all but a few exceptional cases, the Li-salt free Wittig reaction proceeds irreversibly. Below is presented the evidence that has been put forth implying the operation of kinetic control in the reactions of each of the three major classes of ylide. The exceptional cases in which thermodynamic control has been shown to operate in Li-salt free reactions will also be described.

(i) Non-stabilised ylides

It is generally accepted that OPA must occur at some point along the reaction coordinate from phosphonium ylide and aldehyde to alkene and phosphine oxide – either as an intermediate or TS. OPAs have been shown to be the only observable intermediates many times in reactions of non-stabilised ylides. 38,65,66,69,76 It has also been demonstrated in such reactions that the final step of the Wittig reaction is decomposition of OPA to alkene and phosphine oxide, since the rate of OPA decomposition equals the rate of alkene formation (i.e. this is a first order process), 69,70 and since the diastereomeric ratio of the OPA generally matches that of the derived alkene. 60,69,70

The operation of kinetic control in a Wittig reaction is established if the kinetic OPA cis/trans ratio (usually determined by low temperature 31P NMR) is identical to the Z/E ratio of the alkene produced by warming of the OPA. There are surprisingly few examples in which kinetic control has been absolutely proven in this way. Many of the examples of the application of low temperature 31P NMR monitoring of the OPA intermediate of a Wittig reaction have been used to demonstrate the operation of OPA equilibration in reactions involving aromatic or tertiary aldehydes and/or trialkylphosphonium alkylides (see discussion later). However, stereospecific conversion of the OPA distereomers produced in the Wittig reactions shown in Table 1 to alkene has been demonstrated by low temperature 31P NMR, implying the operation of kinetic control in these reactions. 24,38,59,69,70,92,93

As shown in Table 1 entry 1, stereospecific conversion of the OPA (cis/trans ratio of 5:8:1) produced in the reaction of hexanal with n-butylidenetriphenylphosphorane (generated using LiHMDS) to a 5.8:1 mixture of Z and E dec-4-ene has been reported. 69,70 This is consistent with the fact that LiBr has been shown to exert an effect on the stereoselectivity of OPA formation in the reaction of Ph3P=CHCH and PhC6H5CH=CHO, but not to affect the stereochemical ratio of the OPA formed from these reactants if added to a pre-formed solution of the OPA to give betaine-LiBr complex. 65 These observations suggest that OPA formation is irreversible in reactions of aliphatic aldehydes, even in the presence of Li+, and that the diminished Z-selectivity in these reactions arises from the effect of Li+ on the initial formation of OPA.
cis/trans and alkene Z/E ratios. X = halide counter-ion.

<table>
<thead>
<tr>
<th>Entry</th>
<th>Y</th>
<th>Z</th>
<th>R²</th>
<th>R¹ᵃ</th>
<th>R¹ᵇ</th>
<th>Base</th>
<th>OPA cis/trans ratio</th>
<th>Alkene Z/E ratio</th>
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<td>Ph</td>
<td>n-Pr</td>
<td>n-C₅H₁₁</td>
<td>H</td>
<td>LiHMDS</td>
<td>5.8:1</td>
<td>5.8:1</td>
<td>69</td>
</tr>
<tr>
<td>2</td>
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<td>n-Pr</td>
<td>Ph</td>
<td>H</td>
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<td>92</td>
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<td>Ph</td>
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<td>H</td>
<td>NaHMDS</td>
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<td>96:4</td>
<td>69</td>
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<td>H</td>
<td>NaNH₂</td>
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<td>5:95</td>
<td>59</td>
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<td>9:90</td>
<td>93</td>
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<td>Me</td>
<td>Ph</td>
<td>H</td>
<td>NaHMDS</td>
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<td>94:6</td>
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<td>Ph</td>
<td>H</td>
<td>NaHMDS</td>
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<td>96:4</td>
<td>24</td>
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<tr>
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<td>Ph</td>
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<td>NaHMDS</td>
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<td>98:2</td>
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<td>H</td>
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<td>Ph</td>
<td>i-Pr</td>
<td>2-BrC₆H₄</td>
<td>H</td>
<td>NaHMDS</td>
<td>94:6</td>
<td>91:9</td>
<td>38</td>
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<tr>
<td>13</td>
<td>DBP</td>
<td>Ph</td>
<td>i-Pr</td>
<td>c-C₅H₈</td>
<td>H</td>
<td>NaNHMDS</td>
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<td>43:57</td>
<td>38</td>
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<td>14</td>
<td>DBP</td>
<td>Ph</td>
<td>i-Pr</td>
<td>c-C₅H₈</td>
<td>H</td>
<td>NaNHMDS</td>
<td>94.6</td>
<td>91:9</td>
<td>38</td>
</tr>
</tbody>
</table>

Phosphonium salt counter-ion is bromide in all cases except entries 5–7, for which it is iodide. The decomposition temperature is the temperature at which OPA decomposition to alkene was effected experimentally. Unless otherwise specified, decomposition was carried out at “room temperature” (15-20 °C).

⁶ Only one OPA signal observed in ³¹P NMR spectrum obtained at -20 °C.

⁷ The paper in question states that the reaction “did not exhibit stereochemical drift”, but does not quote the alkene Z/E ratio.

⁸ The specific ylide used in this reaction, incorporating the dibenzophosphole (DBP) system, is shown in Fig. 5(a). The experimental OPA decomposition temperature for this reaction was 110 °C. See reference 59.

⁹ Alkene Z/E ratio not reported.

¹⁰ OPA decomposition temperature not specified.

¹¹ “2-fur” indicates a 2-fur-2-yl group. OPA decomposition in this reaction carried out in refluxing THF, quoted as 70 °C.

¹² In a separate reaction, n-BuLi was used to generate the ylide. Stereospecific OPA decomposition was also noted from the OPA in this case, but with the exception of entry 8 the yield of alkene was much lower than when NaHMDS was employed.

¹³ OPA decomposition in this reaction carried out in refluxing THF (heating bath temperature 80 °C).

OPAs are generally air and temperature sensitive. The decomposition products from exposure to (relatively) high temperatures are alkene and phosphine oxide, so at low temperature the OPAs derived from non-stabilised ylides are kinetically stable. If the phosphorus of the OPA is constrained by being in a five-membered ring, as in dibenzophosphole (DBP) derived OPAs, then the OPA is kinetically stable at room temperature, and requires heating to induce alkene formation.⁵⁹ Such OPAs can conveniently be monitored spectroscopically at room temperature, and as a consequence they have been used in a number of studies to elucidate details of the Wittig reaction mechanism.⁵⁹,⁶⁸ DBP-derived OPAs have been generated from Wittig reactions (e.g. from the ylide shown in Fig. 5(a)) and by β-HPS depromotion (e.g. of the β-HPS shown in Fig. 5(b)). They have also been used in similar investigations on reactions of semi-stabilised ylides (see section 4.1 part (ii)).

As shown in Table 2, individual OPA isomers generated by a process independent of the Wittig reaction (by depromotion of β-HPS using sodium or potassium base, which transiently gives betaine and then OPA) have been demonstrated by low temperature ¹H and ³¹P NMR to undergo stereospecific decomposition to alkene. That is, erythro β-HPS gave only cis-OPA and hence only Z-alkene, and likewise threo β-HPS gave trans-OPA and hence E-alkene.⁶⁸,⁶⁹,⁹⁵ Other examples exist in which the β-HPS depromotion technique has been used to demonstrate the operation of kinetic control in Wittig reactions of non-stabilised ylides.⁹⁴

Another method for OPA generation independent of a Wittig reaction involves Sn₂₂ epoxide ring-opening by lithium diphenylphosphide to give β-oxidophosphine, which is quaternised with methyl iodide to give betaine and hence OPA. cis-Epoxy was shown to give E-alkene stereospecifically by this method via trans-OPA, while trans-epoxide stereospecifically gave Z-alkene via cis-OPA.⁹⁵,⁹⁶ The latter is shown in Scheme 14.
The observation of the "stereochemical drift" in all cases of stereochemical drift so far reported, the OPA diastereomeric ratio is a consequence of its greater stability, since it affects the stereoselectivity of OPA formation in reactions of alkene phosphine oxide is stereospecific, irreversible, and occurs through a syn-cycloreversion. Therefore the alkene Z/E ratio is identical to the kinetic OPA cis/trans ratio, and the stereoselectivity of the alkene product is decided in a reaction leading from starting materials to OPA.

There exist some exceptional cases where a non-correspondence has been observed between the diastereomeric ratios of the OPA and of the alkene – this has been termed "stereochemical drift". In all cases of stereochemical drift so far observed, the E-isomer has been favoured to a greater extent in the alkene diastereomeric ratio than the trans-isomer had been in the OPA cis/trans ratio. The augmentation of the production of E-alkene cannot be a consequence of its greater stability, since OPA decomposition is irreversible, and so it must be due to preferential (and probably irreversible) formation of the more thermodynamically favoured of the OPA isomers, the trans-OPA.

Stereocchemical drift is in certain circumstances linked to the presence of Li⁺. As was mentioned above, the presence of Li⁺ affects the stereoselectivity of OPA formation in reactions of alkylidenetriphenylphosphoranes with aliphatic aldehydes (higher [Li⁺] results in more trans-OPA65), but these reactions are nonetheless irreversible.65,66 Similarly, the OPA diastereomeric ratio in the reaction of butyldienetriphenylphosphorane with benzaldehyde is different in the presence of Li⁺ (favouring trans-OPA) than it is in Li salt-free conditions, but the initial ratio

**Table 2.** β-Hydroxyphosphonium salts for which stereospecific decomposition to alkene has been demonstrated for either both isomers or the erythro isomer only.

<table>
<thead>
<tr>
<th>Entry</th>
<th>Y</th>
<th>Z</th>
<th>R1a</th>
<th>R1b</th>
<th>OPA decomposition temperature°C</th>
<th>Ref</th>
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<tr>
<td>1</td>
<td>Ph</td>
<td>Et</td>
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<td>C(Me2)CH2Ph</td>
<td>H</td>
<td>20</td>
</tr>
<tr>
<td>2</td>
<td>DBP</td>
<td>Et</td>
<td>Me</td>
<td>C(Me2)CH2Ph</td>
<td>H</td>
<td>20</td>
</tr>
<tr>
<td>3</td>
<td>Ph</td>
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<td>Me</td>
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<td>Et</td>
<td>Me</td>
<td>c-C6H11</td>
<td>Me</td>
<td>&gt;-50</td>
</tr>
</tbody>
</table>

* Temperature at which OPA decomposition was effected experimentally.

* The specific β-HPS used in this reaction, incorporating the dibenzophosphate (DBP) system, is shown in Fig. 5(b).

* Deprotonation carried out using NaHMDS.

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**Scheme 14:** Reaction system for which OPA is produced by nucleophilic cleavage of an epoxide by diphenylphosphide and quaternisation of the resulting phosphine and shown to decompose stereospecifically to alkene.65,66

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**Figure 5.** (a) (Ethylidene)ethylidenphosphonium,69 (b) erythro or threo (3-hydroxy-4,4-dimethyl-5-phenylpent-2-yl)ethylidenphosphonium iodide, incorporating the dibenzophosphate (DBP) system.69
remains constant with time at low temperature in both Li⁺-free and Li⁺-present conditions.\textsuperscript{69,77} It is only when the Li⁺-present reaction mixture is warmed up that stereochemical drift occurs, resulting in the observation of a change in the diastereomeric ratio on going from OPA to alkene (favouring E-alkene). The Li salt-free case does not exhibit stereochemical drift (as already stated; see Table 1 entry 1).\textsuperscript{65,69} Positive crossover experiments indicate that the conversion of cis-OPA into trans-OPA in the examples mentioned above of reactions of alkylidenetriphenylphosphoranes with benzaldehyde in the presence of Li⁺ involve OPA reversal to ylide and aldehyde.\textsuperscript{65,69}

Thus, in reactions of non-stabilised ylides with benzaldehyde(s), Li⁺ exerts an effect both on the stereoselectivity of the initial formation of OPA and plays a role in bringing about stereochemical drift in reactions of longer-chain alkylides that would otherwise be under kinetic control. The role of Li⁺ in the process of stereochemical drift may be related to the formation of betaine-lithium halide complexes, which have been isolated from reactions of non-stabilized ylides.\textsuperscript{64,65} There exists one report in which betaine-LiBr complexes have been formed in reactions of a semi-stabilised ylide and undergo Wittig reversal, as judged by positive crossover experiments.\textsuperscript{98}

There is also a small and well-defined set of Wittig reactions – all involving non-stabilised ylides – that exhibit stereochemical drift under Li salt-free conditions (“Li⁺-free stereochemical drift”). Li⁺-free stereochemical drift and has been observed only for the OPAs produced in the reactions of trialkylyphosphoranes with tertiary or aromatic aldehydes,\textsuperscript{68,69,92} and in reactions of aromatic aldehydes with the ethylides of triphenylphosphine.\textsuperscript{38,65} ethylidiphosphine\textsuperscript{38} and P-phenyl-5H-dibenzo phosphole.\textsuperscript{38} In the first two cases, the augmentation of trans-OPA at the expense of the cis-isomer occurred at temperatures below which OPA cycloreversion to alkene and phosphate oxide could not occur, and thus it could be monitored by NMR. These experiments represent the only examples in which stereochemical drift has been observed in action. In the latter three cases, the OPA cis/trans ratio was invariant with time below the temperature at or above which alkene formation could occur, and so stereochemical drift could only be detected to have occurred by comparison of the kinetic OPA cis/trans and alkene Z/E ratios.

An example that illustrates the observation of stereochemical drift is the experiment of Vedejs et al. involving deprotonation of the erythro-β-HPS shown in Fig. 6 at –78 °C. This reaction would be expected to give only the cis-OPA that would be produced in the Wittig reaction of ethylidenetriethylphosphorane and 2,2-dimethyl-3-phenylpropenal. However, the reaction mixture upon warming to –40 °C was observed to consist of a 80:20 mixture of cis and trans OPA.\textsuperscript{54} On warming to –5 °C, the amount of trans-OPA was observed to increase at the expense of the cis-OPA, while a small amount of alkene formation also occurred. Decomposition of OPA to alkene and phosphate oxide was effected at >30 °C giving a final Z/E ratio of 17:83. Some 2,2-dimethyl-3-phenylpropenal was also isolated, indicating reversal to ylide and aldehyde. Reversal was confirmed by the observation of crossover product in an experiment involving deprotonation of the same β-HPS in the presence of chlorobenzaldehyde. At least some of this reversal is likely to be due to reversal of the betaine transiently produced upon β-HPS deprotonation, but OPA reversal must be involved in view of the fact that the OPA cis/trans ratio was observed to change after the passage of some time with an increase in temperature. The three-diastereomer of this β-HPS decomposed stereospecifically to E-alkene.\textsuperscript{68}

Stereochemical drift was shown to operate in the reaction of ethylidenetriphenylphosphorane with benzaldehyde by a series of crossover experiments.\textsuperscript{65} The isomeric ratio of the OPA produced from these reactants remained constant with time below –30 °C. When spectroscopic monitoring of the reactions was carried out below this temperature, no crossover products could be observed. The mechanism of equilibration of the OPA isomers was shown conclusively in this case to involve cycloreversion to Wittig starting materials. Li-salt free OPA equilibration was also demonstrated for these reactants by an experiment in which deprotonation of a 3:1 mixture of erythro and threo (1-hydroxy-1-phenyl-n-prop-2-yl)triphenylphosphonium bromide gave a 2:1 mixture of Z and E-1-phenylprop-1-ene,\textsuperscript{69} as shown in Scheme 15(a).

Related to the previous example and in the same study, β-HPS erythro-(1-hydroxy-1-phenyl-n-pent-2-yl)triphenylphosphonium bromide (see Scheme 15(b) and Table 2 entry 5) was shown to undergo stereospecific conversion to Z-1-phenylpent-1-ene.\textsuperscript{69} As mentioned above, the OPA produced in the reaction of Scheme 15(b) was also formed with very high cis-selectivity in the Li-salt free Wittig reaction of benzaldehyde + n-butylidenetriphenylphosphorane, and its decomposition proceeded stereospecifically,\textsuperscript{92} or with negligible stereochemical drift (see Scheme 15(c) and Table 1 entries 2–4).\textsuperscript{69} Thus it appears that the reactions of most types of alkylide with benzaldehyde are under kinetic control; only certain ethylides and alklyidenetrialkylphosphoranes show reversibility.
It is noteworthy that all instances of stereochemical drift observed thus far involve the production of increased amounts of trans-OPA at the expense of the cis-OPA, or the production of a greater proportion of E-alkene than had been present of the trans-OPA. In other words, stereochemical drift always results in an increased proportion of trans-OPA, since E-alkene must be produced from trans-OPA, and alkene formation is irreversible. An obvious conclusion is that the trans-OPA is thermodynamically favoured; this is corroborated by the observation that trans-OPAs typically require higher temperatures to effect their decomposition to alkene and phospine oxide than do their cis counterparts.\textsuperscript{68,69,93} Thus, the proportion of Z-alkene produced in a Wittig reaction always represents a lower bound to the proportion of cis-OPA produced in the first step of the reaction. In other words, high selectivity for Z-alkene is directly indicative of kinetic control, and cannot occur “by mistake” (i.e. due to equilibration) as long as the cis-OPA is indeed thermodynamically disfavoured.

(ii) Semi-stabilised ylides

It is not possible, in general, to detect OPAs derived from semi-stabilised ylides by low temperature NMR. Consequently, a kinetic diastereomeric OPA ratio cannot be established to compare with the final alkene Z/E ratio. However, in one case an OPA produced in a Wittig reaction of a semi-stabilised ylide has been reported. As shown in Scheme 16, the reaction of cyclohexanecarboxaldehyde with the phenylidzenophosphate-derived semi-stabilised ylde allyldienephENylidbenzophosphole in THF-d\textsubscript{8} at \textdegree 78 °C gave an OPA (\textsuperscript{31}P NMR δ = -72 ppm) that decomposed above \textdegree 50 °C to 1-cyclohexyl-1,3-butadiene with Z/E = 5.95.\textsuperscript{60} Such a “constrained” OPA (in which the phosphorus atom is a spiro centre between four and five-membered rings) is kinetically relatively stable (compared to other semi-stabilised ylide-derived OPAs\textsuperscript{90}) because increasing bond angle strain is induced in the phosphate ring on going from OPA to phosphine oxide. This effectively raises the barrier to OPA decomposition to alkene and phosphine oxide, and may even lower the barrier to OPA formation.

In addition to the impossibility of spectroscopic observation of OPAs in Wittig reactions of semi-stabilised ylides, with the above singular exception, crossover experiments with such ylides\textsuperscript{9} do not give meaningful results because the reaction of the semi-stabilised ylide proceeds so quickly that the reaction is already complete by the time the crossover ylide or aldehyde is added. Thus, proof for the operation of kinetic control in Wittig reactions of semi-stabilised ylides requires OPA generation from a source other than a Wittig reaction, and demonstration of diastereospecific conversion either (a) of the precursor of cis or trans-OPA to Z or E-alkene, respectively, or (b) of a mixture of the precursors of known diastereomeric ratio to alkene of corresponding Z/E ratio. This has been done for reactions of semi-stabilised ylides by forming alkene from either (i) β-HPS by deprotonation or (ii) epoxide by nucleophilic ring cleavage by phospide followed by protonation of the alkoxide and then quaternisation of phosphorus. Both methods presumably form betaine and OPA transiently en route to alkene, and have both been described in connection with reactions of non-stabilised ylides in section 4.1 part(i).

Each of these methods suffers from the fact that betaine generation is inherent in the process of OPA formation. The formation of OPA in Wittig reactions does not necessarily require the intermediacy of a betaine. There may exist processes in experiments involving β-HPS deprotonation or nucleophilic cleavage of epoxide that do not necessarily operate in Wittig reactions which result in the OPA cis/trans ratio being different to the diastereomeric ratio of the starting material. For example, dissociation of the transiently formed betaine into ylide & aldehyde (“betaine reversal”), and non-stereospecific recombination of these Wittig starting materials to give OPA may occur in either type of experiment, whereas Wittig reactions do not necessarily form betaines. Also in β-HPS deprotonation experiments, there exists the possibility of the formation of β-hydroxy ylide (by deprotonation of the α-carbon, not the β-hydroxy group), which then undergoes non-stereospecific proton transfer to form betaine and hence OPA. This latter process has been demonstrated to occur in corresponding experiments involving β-HPSs derived from stabilised ylides when strong bases are used (see section 4.1 part (iii)).\textsuperscript{60} In each case, stereospecific decomposition of the OPA mixture of changed diastereomeric ratio to alkene product with a Z/E ratio that is not identical to the initial β-HPS erythro/threo ratio or epoxide diastereomeric ratio could be erroneously interpreted as indicating the operation of some equilibration of OPA diastereomers in the corresponding Wittig reaction. The same problem of the apparent operation of stereochemical drift applies in crossover experiments involving β-HPS deprotonation.

Despite such limitations, these methods have been successfully used to demonstrate the irreversibility of OPA formation in reactions of semi-stabilised ylides.

The cis and trans-OPAs that would be produced in the reaction of allyldienemethylidbenzophosphole with cyclohexanecarboxaldehyde were independently synthesised in separate experiments by deprotonation of erythro and three β-HPSs \textsuperscript{11} and \textsuperscript{12} respectively (see Chart 2).\textsuperscript{60} Similarly, the cis and trans-OPAs from the reaction of benzyldeinemethylidbenzophosphole with cyclohexanecarboxaldehyde were obtained by deprotonation of β-HPSs \textsuperscript{13} and \textsuperscript{14} respectively (see Chart 2). Each deprotonation experiment gave a single OPA (as evaluated by low temperature \textsuperscript{31}P NMR), which decomposed stereospecifically to the expected alkene isomer – i.e. erythro β-HPS gave Z-alkene, and three-β-HPS gave E-alkene. These experiments prove that there is no equilibration of either the betaine or of the OPA which are each necessarily formed in the path from β-HPS to alkene. The corresponding Wittig reaction must occur under kinetic control, since the formation of each of the possible intermediates is irreversible.
Kinetic control has also been shown indirectly to be in operation in the reactions of unconstrained semi-stabilised ylides (i.e. where the phosphorus atom of the ylide is not part of a ring system), although as alluded to above, observation of the transient OPA intermediate by NMR is not possible in these cases. Thus, deprotonation of (2-hydroxy-1-phenyl-2-phenylethyl-1-y)lmethyl diphenylphosphonium iodide with each of n-BuLi, NaHMDS and KHMDS gave Z-alkene stereospecifically. Treatment of stilbene oxide with lithium diphenylphosphide gave a β-oxidophosphine intermediate, which was quaternised with methyl iodide, resulting in the formation of alkene, presumably by transient formation of a betaine which then underwent ring-closure to OPA as shown in Scheme 17. The process was shown to proceed with inversion – trans-stilbene oxide gave Z-stilbene, and cis-stilbene oxide gave E-stilbene. These results show that there is no inter-conversion of OPA intermediates derived from semi-stabilised ylides that have been generated by non-Wittig processes (and also that there is no interconversion of the associated betaines, whether they are involved in the Wittig mechanism or not). This confirms that both OPA formation and OPA decomposition to alkene and phosphate oxide are irreversible, at least in the case of Wittig reactions of P-phenyl-5H-dibenzophosphet-derived semi-stabilised ylides, and also most significantly in the case of Wittig reactions of unconstrained methyldiphenylphosphine-derived ylides with benzaldehyde.

(iii) Ester-stabilised ylides
No OPAs derived from stabilised ylides have ever been detected by NMR, not even dibenzophosphole derived ones. Consequently, proof for the operation of kinetic control in reactions of these ylides rests on the demonstration of stereospecific conversion of erythro or threo β-HPS to Z or E alkene, respectively. The erythro-β-HPSs derived from primary, secondary, tertiary, and aromatic aldehydes and ester-stabilised ylide (ethoxy carbonylmethylidene)methyl diphenylphosphorane were obtained by addition of phosphino-enolate [Ph2PCHCOOEt]Li to aldehyde followed by acid quenching to

Scheme 17. Stereospecific conversion of (a) cis-stilbene oxide to E-stilbene and (b) trans-stilbene oxide to Z-stilbene by nucleophilic ring cleavage and quaternisation of the resulting β-oxidophosphine.

Scheme 18. Synthesis of β-hydroxyphosphine, and hence β-HPS corresponding to the acid quench product of a Wittig reaction of a stabilised ylide, and stereospecific conversion of the erythro-β-HPS to Z-alkene. R = Ph, CH2=CHPh, c-C6H11, or CMes=CH2Ph. Deprotonation experiments using the β-HPSs thus obtained were carried out in THF and ethanol. DBU, 1,8-diazabicyclo[5.4.0]undec-7-ene, mesityllithium and KHMDS bases were used in separate experiments. Two equivalents of p-chlorobenzaldehyde were added to the reaction mixture 30 seconds after the addition of base. This aldehyde reacts with any ylide produced by potential reversal of OPA or betaine to Wittig starting materials, thus preventing recombination of the reactants derived from reversal. Delayed addition of p-chlorobenzaldehyde prevented it from exerting any effect on the betaine that is necessarily transiently produced in β-HPS deprotonation. High yield and stereospecificity and negligible crossover were observed in the conversion of the erythro-β-HPSs to Z-alkene (Z/E > 98:2) using DBU as base for the cases with R = CH2=CHPh, c-C6H11, or CMes=CH2Ph. Greater production of each of crossover alkene and E-cinnamate was observed in the case with R = Ph. In deprotonation experiments at 20 °C using KHMDS in THF, the production of crossover product was observed to have increased relative to the experiments with DBU – 8% for the case with R = CH2=CHPh, 5% for R = c-C6H11, 23% for R = CMes=CH2Ph and 48% for R = Ph. The production of crossover alkene showed that ylide derived from some form of reversal to Wittig starting materials had to have been present. E-Cinnamate could not have resulted from any OPA or betaine reversal process since ylide would have been intercepted as crossover product, and it was shown not to result from isomerisation of the Z-cinnamate under the reaction conditions.

The pathway by which crossover product and E-alkene were formed was elucidated using α-deuterated erythro-β-HPSs (see Scheme 19). Under “strong base conditions”, deprotonation of α-
deuterated *erythro*-β-HPS at low temperature (-78 °C) yielded a very high proportion of Z-alkene with high deuterium content, while at higher temperatures a much lower proportion of Z-alkene was produced, and its deuterium content was also much lower. Most significantly, no deuterium could be detected under any circumstances in the E-alkene or crossover alkene. None of these observations are consistent with reversal of the betaine or OPA produced by β-HPS deprotonation, because if this was undergoing reversal then the E-alkene or crossover alkene should show some deuterium content.

In explanation of the experimental observations (see Scheme 19), it was proposed that the base can remove either the hydroxyl proton (path A) or the α-deuteron (path B). In the former case, the betaine formed reacts to give cis-OPA and hence Z-cinnamate with full deuterium content. In the latter case, a β-hydroxy ylide (that has lost its deuterium) results, which can undergo proton transfer (by an unknown mechanism) to either face of the β-hydroxy ylide yielding *erythro* or *threo* betaine (the latter being favoured). This can either undergo ring closure to give cis or trans-OPA and hence deuterium-free Z- or E-alkene, respectively, or can revert to ylide and aldehyde and hence to crossover alkene. Dedeuteration is more likely at higher temperature (due to poorer discrimination of acidic sites by the base), so that more of each of the crossover product and the E-alkene are observed. There is also a lower proportion of deuterium in the Z-alkene that is produced, as some of it is derived from non-deuterated *erythro*-betaine, which is itself derived from β-hydroxy ylide. The fact that the Z-alkene alone retains the deuterium label in these experiments confirms that cis-OPA is formed irreversibly and that it undergoes stereospecific decomposition to Z-alkene. This means that OPA is necessarily the final intermediate in the Wittig reaction of stabilised ylides, regardless of how it is formed.

That crossover alkene is produced in these deprotonation experiments (especially the ones involving strong base) indicates that it is highly likely that some ylide is present after deprotonation has occurred. It has been explained above how it was shown that this ylide cannot be derived from OPA. If formed, it must thus result from C-C bond scission in either the betaine or the β-hydroxy ylide. There is no definitive evidence to prove which entity undergoes reversal to ylide and aldehyde.

Betaine reversal is not an issue for the *erythro*-betaine produced by hydroxyl deprotonation of the initial *erythro*-β-HPS, as proved by the complete absence of deuterium labelled E-enoate in all cases, but based on the above results, it is in principle possible. If betaine reversal does occur, a possible explanation given by Vedejs and Fleck is that β-HPS deprotonation by weak base (e.g. DBU) produces a syn-betaine rotamer that can easily cyclise to OPA, whereas proton transfer involving a β-hydroxy ylide or β-HPS deprotonation by strong base (e.g. KHMDS) may produce rotamers that are more prone to reversal to Wittig starting materials.

It may be assumed that betaine (whether produced as syn or *anti* rotamer) is energetically uphill from Wittig starting materials based on its apparent readiness to decompose to ylide and aldehyde or to alkene and phosphine oxide via OPA. Certainly OPAs produced in Wittig reactions of non-stabilised and semi-stabilised ylides have been demonstrated to be the only observable intermediates in those reactions, so in these cases, OPA is more stable than betaine. Since *cis*-OPA is shown by the above experiments to be formed irreversibly, and to decompose very quickly to alkene and phosphine oxide, *E*-selectivity in reactions of stabilised ylides cannot be the result of the decomposition of *trans*-OPA being more rapid than that of *cis*-OPA. The remaining possibility for involvement of betaine in reactions of stabilised ylides requires that they equilibrate *via* ylide and aldehyde, and that ring closure for *threo*-betaine is kinetically preferred compared to that for the *erythro* isomer. This would require there to be a large energy barrier between betaine and OPA. In light of the rapidity of alkene formation in β-HPS deprotonation experiments, and of the fact that the first step of Wittig reactions of stabilised ylides (the bimolecular step) is rate-limiting, this seems highly unlikely. All signs indicate the operation of kinetic control in reactions of stabilised ylides. Since the betaine mechanism cannot account for high *E*-selectivity under kinetic control, it can be concluded that it is not in operation in reactions of stabilised ylides.

Another (perhaps less likely) possibility for the formation of crossover product is that the β-hydroxy ylide itself perhaps might undergo a Wittig reaction with the *p*-chlorobenzaldehyde, giving the OPA shown in Scheme 20. This OPA can give crossover alkene either directly by elimination of the non-crossover aldehyde, or by forming β'-hydroxy enoate, which could itself eliminate aldehyde.

Since the β-hydroxy ylide pathway cannot be involved in a normal Wittig reaction, and since it plays at most a minor role in stabilised ylide-derived β-HPS deprotonation experiments in the absence of strong base, it may be concluded that *erythro*-β-HPSs derived from stabilised ylides decompose stereospecifically to Z-alkenes upon deprotonation. Thus, the typically observed high *E*-selectivity in Wittig reactions of ester-stabilised ylides results from a kinetic preference for the formation of *trans*-OPA (i.e. the pathway leading to E-alkene), since OPA decomposition to alkene & phosphine oxide is shown by these results to be
stereospecific, and is in general irreversible. Irreversible formation of trans-OPA in reactions of stabilised ylides can be inferred from the fact that cis-OPA is formed irreversibly, since trans-OPAs are generally thermodynamically favoured over cis-OPA and decompose more slowly than cis-OPAs (with few exceptions), indicating a greater barrier to cycloreversion for trans-OPAs.

(iv) Further evidence for kinetic control in Wittig reactions of all ylide types

Based on the material presented in sections (i) to (iii) above, it can be concluded that any Wittig reaction that forms predominantly Z-alkene is under kinetic control, regardless of whether irreversibility of OPA formation has been explicitly demonstrated. There are many examples in the review of Vedejs and Peterson of highly Z-selective Wittig reactions which are thus irreversible.

Very recently, we have shown that reactions of both aromatic and aliphatic aldehydes bearing a β-heteroatom substituent systematically show very high selectivity for Z-alkene (or its precursor, cis-OPA) in their reactions with non-stabilised, semi-stabilised and ester-stabilised ylides. Control reactions of analogous aldehydes lacking the suitably-placed heteroatom with the same ylides show much lower selectivity for Z-alkene, and in fact in many cases these reactions are E-selective. The existence of this counter-intuitive effect that is common to all three ylide types indicates that not only are the reactions of these ylides irreversible, but, more significantly, that they have a common mechanism.

That the results involving the β-heteroatom-substituted aldehydes mentioned above is indicative of general irreversibility of OPA formation in the Wittig reaction of all ylide types may be demonstrated by the following argument: Let us assume that any E-selective Wittig reaction is so as a consequence of reversible formation of cis-OPA. The activation barrier for cycloreversion of cis-OPA to ylide + aldehyde (ΔG‡1 in Fig. 7) must therefore be surmountable. In the reactions of the β-heteroatom substituted aldehydes, the free energy of the reactants and OPA intermediates is similar to that of the analogous species in reactions where the aldehyde lacks a heteroatom substituent. If it is correct, as seems likely, to assume that the cis-OPA is thermodynamically disfavoured relative to the trans-isomer in these reactions, then the energy of the cis-selective TS on the path from reactants to OPA must be lower than in the E-selective reactions. Therefore the activation barrier for cycloreversion of cis-OPA to ylide + aldehyde in reactions of β-heteroatom substituted aldehydes (ΔG‡2 in Fig. 7) is lower than in the E-selective reactions. If OPA reversal was a significant factor in the stereoselectivity of the E-selective reactions, then reactions of β-heteroatom substituted aldehydes should also be E-selective, but to an even greater extent.

Fig. 7. Reaction coordinate diagrams for formation of cis-OPA in Wittig reactions of stabilised ylides, with diagrams of TS geometries.

4.2. The nature of the first formed intermediate in the Wittig reaction

The non-involvement of betaine intermediates (or spin-paired diradical C-C bonded species proposed by McEwen and co-workers) in Wittig reactions of non-stabilised ylides has been conclusively demonstrated using a phenyldibenzophosphole-derived ylide, iso-propyldi phenyldibenzenophosphole (see Schemes 21 & 22). The OPA derived from this ylide can be detected by NMR. It is pseudorotationally restricted because of the high ring strain that would be induced if the five-membered ring were forced to span two equatorial sites. The two possible OPA pseudorotamers that have the five-membered ring spanning axial and equatorial sites interconvert sufficiently slowly at low temperature that they can be distinguished by 31P NMR. Their equilibrium pseudorotameric ratio was established by observation of the 31P NMR of the Wittig reaction of iso-propyldi phenyldibenzenophosphole and 3-phenylpropanal at –78 °C in Et2O (see Scheme 21) to be 1.8:1. The kinetic OPA pseudorotameric ratio (initial ratio) for this Wittig reaction in Et2O was observed to be 6.5:1 by the addition of the aldehyde to the ylide at –109 °C, and subsequent 31P NMR observation of the reaction mixture at –78 °C.

The kinetic OPA pseudorotameric ratio produced in the deprotonation of the corresponding β-HP in Et2O at –78 °C (see Scheme 22), which must necessarily proceed through a betaine, was also observed by 31P NMR at –78 °C, and was found to be 1.4:2 - different from the kinetic pseudorotameric ratio produced in the Wittig reaction, and also from the equilibrium ratio of the pseudorotamers at this temperature.

The kinetic OPA pseudorotameric ratios were also determined in a similar manner for this reaction at low temperature in THF and CH2Cl2. In each case, the ratio was found to be different for the OPA formed by Wittig reaction compared to that formed by β-HP deprotonation. That a different kinetic pseudorotameric ratio is produced in each instance shows that the formation of OPA by Wittig reaction does not occur through a betaine intermediate, at least in the case of these non-stabilised ylides.

In addition to this very conclusive direct evidence that OPAs are the first-formed and only intermediates in Wittig reactions, the very substantial evidence presented above for the operation of kinetic control in Wittig reactions can itself be interpreted as
### 4.3. Oxaphosphetane structure and cycloreversion mechanism

Westheimer’s rules on apical entry and departure requires that any bond made to phosphorus to form a trigonal bipyramidal species must initially place the new substituent in an apical site, and that upon departure of a substituent from a trigonal bipyramidal species the bond broken must be to a substituent in an apical site.\(^3\) Trigonal bipyramidal species are stabilised to the greatest extent possible if the most electronegative atom(s) occupy the apical sites. This arrangement of the substituents results in the HOMO being as low in energy as is possible.\(^{100}\) Based on these concepts, the most stable OPA pseudorotamers should have oxygen in an apical site, and formation of OPA in the Wittig reaction should (at least initially) give a pseudorotamer of OPA with apical oxygen. Furthermore, decomposition of OPA to alkene and phosphine oxide by breakage of the P-C and C-O bonds must involve an entity in which ring carbon-3 (formerly the ylide α-carbon) is in an apical site and therefore, necessarily, oxygen is in an equatorial site. It has long been thought that this set of constraints necessitated the formation of a second OPA intermediate (with apical C) from the initially formed O-apical OPA by a pseudorotation process.\(^3\) The available computational evidence indeed suggests that the formation of a C-apical OPA is necessary.\(^{42}\) A C-apical OPA should be less stable than one with apical oxygen, for the reason stated above, and certainly no C-apical OPA has ever been observed spectroscopically in the course of a Wittig reaction. Vedejs & Marth have questioned whether the formation of any particular pseudorotamer (or set of pseudorotamers) is essential to the mechanism of decomposition of the initially formed O-apical OPA.\(^{90}\)

Recently, Lamertmsa & co-workers have shown that permutation of ligands in trigonal bipyramidal species may occur in a single step by processes that are comprised of multiple consecutive Berry pseudorotations.\(^{91,101,102}\) González, López-Ortiz & co-workers have shown that one or more of these single step mechanisms is highly likely to operate in the interconversion (stereomutation) of OPAs very similar to those produced in typical Wittig reactions.\(^{40}\) Although there exists at least one example of a Wittig reaction of a chiral ylide in which the chirality of the phosphorus centre is retained in the phosphine oxide product,\(^{79}\) the implication of the results of González, López-Ortiz & co-workers is that retention of configuration at phosphorus is unlikely in any Wittig reaction in which OPA stereomutation is faster than OPA decomposition.\(^{40}\) Of even greater significance is that the same group have shown computationally that OPA cycloreversion may occur in a single step without the need for the formation of a C-apical OPA intermediate. If this is the case, then the C-apical OPA would not be a local minimum on the reaction coordinate, but would still occur on the pathway from O-apical OPA to the cycloreversion TS. In either case, the TS for the cycloreversion process would be a pseudo-trigonal-bipyramidal entity with the (elongated) P-C bond in an apical position. The computational and experimental activation parameters for the decomposition of the OPA were found to agree closely with each other, lending support to the single step mechanism.\(^{40}\)

Below is presented the experimental data that has been amassed pertaining to the structure of OPAs produced in Wittig reactions and to OPA pseudorotation.

Vedejs and Marth were able to resolve signals due to two different pseudorotamers of a dibenzophosphole-derived OPA (see Chart 3 structures 16 and 17) by low temperature \(^1\)H NMR, and to prove that these pseudorotamers had oxygen in the apical position.\(^{90}\) This was established by an evaluation of the one bond coupling \(^{31}\)C-\(^31\)P constants \(J_{PC}\) in the \(^{13}\)C NMR for the phosphorus-bound carbons in the dibenzophosphole unit. In such a system there must necessarily be one carbon apical and one equatorial. Indeed, by low temperature \(^{13}\)C NMR, there appear three signals for quaternary aromatic carbons with small coupling constants to phosphorus (less than 20 Hz) – diagnostic of aromatic carbons in an apical position or not directly joined to phosphorus and thus not in the trigonal bipyramid – and a fourth quaternary aromatic carbon showing coupling to phosphorus of 132 Hz. The latter is characteristic of an sp\(^2\)-hybridised equatorial carbon in a phosphorus-centred trigonal bipyramid.\(^{103}\) The OPA ring carbon (\(\delta 54.5\)) decisively shows \(J_{PC} = 82\) Hz, which is indicative of an equatorial aliphatic (sp\(^3\)) carbon in a phosphorus-centred trigonal bipyramid. This means that the most stable OPA in solution has either the structure 16 or 17 in Chart 3, but not structure 18. In a separate publication, Vedejs and Marth disclosed the resolution by low temperature \(^1\)H and \(^{31}\)P NMR of signals due to two different pseudorotamers of another DBP-derived OPA (shown above in Scheme 21),\(^{74}\) for which the presence of an apical oxygen was established. This data was used by Vedejs & Marth,\(^{74}\) and by ourselves,\(^{38}\) to establish that other dibenzophosphole-derived OPAs (19-23 in Chart 3) as well as some examples of unconstrained OPAs are also likely to have...
Oxygen in the apical position in their favoured solution pseudorotamer(s). Other spectra collected on unconstrained OPAs with the ylide α-carbon \(^{13}\)C labelled show \(^1J_{PC}\) values of very similar magnitude for carbon-3 in the OPA ring.  

Bangerter et al. resolved signals due to pseudorotamer OPAs 24a and 24b (see Chart 3), and in a separate experiment those of the same for OPA pseudorotamers 25a and 25b by low temperature \(^{31}\)P NMR. They established a \(^1J_{PC}\) value of 86.7 Hz for ring carbon-3 in OPAs 25a and 25b, thus showing that the dominant solution structure for these OPAs has oxygen in an apical site. The \(^1J_{PC}\) values for the \(P\)-ethyl group of 25a and 25b labelled with \(^{13}\)C at the methylene group were determined to be 92.4 and 106.8 Hz respectively. These values are characteristic of an equatorial sp\(^3\)-hybridised group, which leads to the conclusion that the alkyl group is preferentially placed in an equatorial position in both pseudorotamers.

In the studies of Vedejs and Marth, the rate of pseudorotation of each OPA was established by line shape analysis of variable temperature NMR spectra of the resolved OPA pseudorotamers. It was concluded that in each case the rate of OPA pseudorotation is too fast relative to the rate of OPA decomposition to have any effect on the decomposition rate. The pseudorotameric forms of dibenzophosphole-derived OPAs that have the P-O bond and one of the 5-membered ring P-C bonds in the apical positions of the phosphorus-centred trigonal bipyramid are very strongly favoured thermodynamically over other possible pseudorotamers. There are only two pseudorotamers of any dibenzophosphole-derived OPA that satisfy these constraints on the positioning of the phosphorus-substituents, and hence a maximum of two major signals only should be expected (and are observed) in the \(^{31}\)P NMR spectra of these OPAs. For example, pseudorotamers 19 and 20 are the two possibilities for the OPA produced in the reaction of (ethylidene)ethylidibenzophosphole with acetone (see Chart 3). The \(^{13}\)C NMR spectrum of this OPA (19/20) contains only one set of signals. This can mean either that 19 and 20 are undergoing rapid interconversion in solution or that one of these two OPA pseudorotamers is heavily favoured in solution. Other OPAs (21a/22a and 21b/22b, see Chart 3) also only show a single set of signals by \(^{13}\)C NMR. Since the pseudorotamers of OPAs very similar to these can be resolved by low temperature NMR, Bangert et al. concluded that in the case of each of the OPAs 19/20, 21a/22a and 21b/22b, one of the two possible pseudorotamers is predominant in solution, and that the pseudorotamers do not interconvert rapidly at low temperature.

### Chart 3. Dibenzophosphole-derived OPAs 16-23, and ethylidiphenylphosphine-derived OPAs 24 and 25

4.4. How does the observed selectivity for Z or E alkene in the Wittig reaction arise? An explanation of how the currently accepted mechanism explains the stereoselectivity.

- **(i)** Summary of experimental evidence:
  - (a) OPAs are the only observable intermediates in reactions of non-stabilised ylides.
  - (b) Betaines and diradical species have been proven not to play a part in the reactions of these ylides.
  - (c) Stereoselective conversions of erythro and threo betaines that would be derived from semi-stabilised ylides (generated either from β-HPs or epoxides) to Z and E alkene respectively have been demonstrated.
  - (d) Stereoselective conversion of dibenzophosphole-derived erythro-β-HPS to cis-OPA (which would be formed from semi-stabilised ylide) by deprotonation (presumably via erythro-betaine) and hence Z-alkene has been observed by low temperature NMR. This has also been done for the conversion of threo-β-HPS to E-alkene. A DBP-derived OPA has also been observed in the case of one Wittig reaction of a semi-stabilised ylide, and was demonstrated to undergo stereospecific decomposition to alkene and phosphine oxide.
  - (e) Stereoselective conversion of erythro-betaines derived from stabilised ylides (and generated from erythro-β-HPs) to Z-alkenes has also been observed.
  - (f) It seems safe to assume that conversion of trans-OPA derived from stabilised ylides to E-alkene is irreversible and stereospecific, since E-alkene formation is generally very favoured in such reactions, and also trans-OPAs are typically more stable than their cis counterparts.
  - (g) OPAs are obligatory at some point along the Wittig reaction coordinate (either as intermediates or TSs). Decomposition of OPA derived from non-stabilised ylides has been shown to occur at the same rate as alkene formation. OPAs derived from all three types of ylide has been shown to decompose irreversibly and stereospecifically. Thus it has been shown that OPA is the final intermediate in the Wittig reaction (regardless of how it is formed).
  - (h) The betaine mechanism cannot account for kinetic E-selectivity.
  - (i) These facts, all together, strongly suggest that OPAs are the first-formed and only intermediates formed in Li-salt free Wittig reactions.
(ii) Selectivity in reactions of non-stabilised ylides

By the convention employed in the paper of Aggarwal, Harvey and coworkers,\(^\text{42}\) when the ylide P=C bond and the aldehyde C=O bond are parallel (i.e. a planar TS), the puckering angle of the TS is defined as 0° (see Fig. 8(a)). If the dihedral angle between the carbonyl C=O bond and the ylide C-R\(^2\) bond is smaller than the corresponding angle in the planar TS (see Fig. 8(b)), then the TS puckering angle is positive. If the dihedral angle is greater than in the planar TS, then the TS puckering angle is negative (see Fig. 8(c)).

The following is Vedejs’s rationale for the mechanism of the Wittig reaction of non-stabilised ylides.\(^\text{49}\) Alkylideneri phenylphosphoranes (non-stabilised ylides) react preferentially through an early, puckered, transition state in which the carbonyl substituent occupies a pseudo-equatorial position (see Fig. 9(a)). The forming P-O and C-C bonds are long and rehybridisation about the ring atoms is not particularly advanced.

In the lowest energy TS, the puckering angle is positive. This arrangement minimises steric interactions between the carbonyl substituent and the dimethylphenyl groups (referred to as “1-3 interactions”, with the numbering of the ring positions as shown in Fig. 9), which are still in a pseudo-tetrahedral arrangement about phosphorus. It also results in the minimisation of steric interactions between the carbonyl substituent and the ylide substituent (“1-2 interactions”), and allows the forming P-O bond to avoid the P-phenyl group that is necessarily projecting in the direction of carbonyl approach to the ylide. The steric interactions of the ylidic substituent R\(^2\) with the carbonyl substituent (“1-2 interactions”) and with the P-phenyl groups (“2-3 interactions”) are minimised if it is in a pseudo-axial site, as shown in Fig. 9(a), and hence this TS leads to cis-OPA and Z-alkene. Calculations on the Wittig reaction of Ph\(_3\)P=CHMe and MeCHO at the B3LYP/6-31G\(^*\) (THF) level of theory (employing a continuum solvent with the dielectric properties of THF) confirm this to be the lowest energy TS, and indicate that there may be a stabilising hydrogen bonding interaction between the aldehyde oxygen and one of the P-phenyl C-H bonds in both diastereomeric TSs.\(^\text{42}\)

In contrast to the cis-selective TS, there is no clear picture of the precise geometry of the lowest energy trans-selective TS in reactions of alkylideneri phenylphosphoranes. A possible puckered trans-selective TS with the ylide \(\alpha\)-carbon substituent in a pseudo-equatorial site (Fig. 9(a) with R\(^2\) and H swapped at carbon-2) would suffer from large 1-2 interactions and hence would be disfavoured relative to the cis-selective TS. A possible planar trans-selective TS is shown in Fig. 9(b). Like the cis-selective TS described above, bond formation and rehybridisation about the ring atoms are each not particularly advanced - so the TS is early. 1-2 steric interactions between the carbonyl substituent and the ylide \(\alpha\)-carbon substituent are avoided in this TS due to the trans arrangement of these substituents. However, it is less stable than the cis-selective TS because one of the P-phenyl groups encumbers the approach of the carbonyl oxygen to phosphorus, and because of large 1-3 steric interactions between the carbonyl substituent R\(^1\) and one of the P-phenyl groups. As with the cis-selective TS there may exist hydrogen bonding between the carbonyl oxygen and a P-phenyl C-H in this TS.\(^\text{42}\)

It is postulated that the OPAs form irreversibly (as has been very comprehensively shown experimentally), and decompose stereospecifically to alkene and phosphine oxide by [2+2] cycloreversion, so stereoselectivity is decided in the cycloaddition step. The computational results of Aggarwal, Harvey & coworkers\(^\text{42}\) on the reaction of Ph\(_3\)P=CHMe + MeCHO indicate that cis-OPA is less thermodynamically stable than the trans-OPA due to increased 1-2 steric interactions in the former, and that OPA formation is exothermic. The barrier to OPA cycloreversion to alkene and phosphine oxide is found to be higher for cis-OPA than for trans-OPA. Based on this, it can be surmised that in some circumstances the barrier to cycloreversion to alkene and phosphine oxide for cis-OPA is similar to the barrier to reversal to ylide and aldehyde. This may provide an explanation for the observation of stereochemical drift in certain reactions of alkylideneri phenylphosphoranes – in particular those with aromatic aldehydes.\(^\text{55,92}\)

Non-stabilised ylides for which one or more of the P-phenyl groups are replaced by alkyl group(s) show much lower Z-selectivity than alkylideneri phenylphosphoranes;\(^\text{4}\) in the context of the cycloaddition mechanism this is explained by decreased 1-3 interactions and thus a to lower propensity to cycloaddition TS puckering. Computational results on the reaction of Me\(_2\)P=CHMe + MeCHO (B3LYP/6-31G\(^*\) (THF)) indicate that both the cis and trans-selective cycloaddition TSs are planar, and very similar in energy – the role of 1-2 steric interactions does not particularly mitigate against the cis-selective TS in this case due to the very long forming bonds in an early TS.\(^\text{42}\) In reactions of alkylideneri alky1phosphoranes with tertiary or aromatic aldehydes, an extra factor (stereochemical drift – see section 4.1 part (i)) is known to be at play in the experimentally observed E-selectivity. The cis-OPA is believed to undergo reversal to ylide and aldehydes, while the trans-OPA does not.\(^\text{68,92}\) and thus trans-OPA accumulates and gives rise to E-alkene. The computational study, referred to earlier, found that the kinetic barriers to cis-
OPA cycloaddition and reversal to ylide and aldehyde were similar, which is consistent with the experimentally observed depletion of this intermediate. The lower Z-selectivity in reactions of alkylidiphenylphosphine-derived ylides compared to alkylidenetriphenylphosphoranes is, however, truly dependent on the energetic discrimination of the TSs in the OPA forming cycloaddition step, as irreversible OPA formation has been proven for such ylides.\textsuperscript{68} Thus in those reactions, the higher E-selectivity reflects an increased kinetic advantage (or decreased kinetic disadvantage) for trans-OPA formation.

A further aspect of the currently accepted mechanism for Wittig reaction of non-stabilised ylides has to do with pseudorotation of the OPA intermediate. The OPA formed by the cycloaddition step is postulated to have the oxygen in the apical position, in accordance with Westheimer’s rule on apical entry and departure for a phosphorus-centred trigonal bipyramid.\textsuperscript{99} This OPA, which could be any one of a number of pseudorotamers with apical oxygen, is proposed to undergo rapid pseudorotation to a less stable pseudorotamer with ring carbon-3 in an apical position, and the ring oxygen in an equatorial site. This pseudorotamer undergoes [2+2] cycloreversion to alkene and phosphine oxide. The computational work of Aggarwal, Harvey & co-workers is consistent with the operation of this process.\textsuperscript{42}

Calculations on the reactions of MeCHO with Ph₃P=CHMe and Me₂P=CHMe respectively show that betaines are significantly higher in energy than OPAs and likewise TSs involving anti addition of the reactants (in the style of an aldol reaction to form an anti-betaine) are higher in energy than cycloaddition TSs.\textsuperscript{42}

- The cycloaddition mechanism of Vedejs for the Li-salt free Wittig reaction of non-stabilised ylides – that being direct, irreversible cycloaddition of ylide and aldehyde to give OPA, followed by facile pseudorotation of the OPA and then by irreversible, stereospecific [2+2] cycloreversion of OPA to alkenyl and phosphine oxide – is consistent with the following facts:
  - OPA formation has been shown to be irreversible,\textsuperscript{68,69,93} and the OPA cis/trans ratio corresponds to the Z/E ratio of the alkene ultimately obtained from the reaction.\textsuperscript{68,69,93} This implies that the stereoselectivity is determined in the OPA forming step. OPA decomposition has been shown to occur at the same rate as alkene formation for reactions of alkylidenetriphenylphosphoranes.\textsuperscript{68,69,70,93} Thus OPA is the final intermediate formed in the Wittig reaction, and phosphine oxide is formed by syn-elimination.
  - Tests for the involvement of radicaloids or ionic species (betaines) indicate that neither type of species play a part in the Wittig reaction.
  - OPA accumulates during the reaction, and OPA decomposition is the final step of the reaction. Therefore this step is rate-determining. It has been observed that reactions of non-stabilised ylides with benzaldehydes have p values for the carbonyl component of 0.2 to 0.59.\textsuperscript{57,104,105} A kinetic isotope effect of 1.053 has also been observed in the Wittig reaction of iso-proplidenetriphenylphosphorane with benzophenone with a \textsuperscript{13}C-labelled carbonyl group.\textsuperscript{57} Thus in the rate-determining step the bonding is changing at the carbonyl carbon. The decomposition of each of the cis and trans OPAs (independently generated form β-HPSs) derived from (ethylidene)ethylidiphenylphosphorane and cyclohexyl methyl ketone has been shown to be first order, and to have a positive entropy of activation.\textsuperscript{93} The results from investigations on the kinetics of reactions of non-stabilised ylides thus support the operation of a rate-determining cycloreversion as the final step of the reaction.
  - Z-selectivity increases in line with the increasing steric bulk of the aldehyde in reactions with a common ylide; so tertiary aldehydes show higher Z-selectivity than secondary aldehydes, which show higher selectivity than primary aldehydes. Likewise, increasing the steric bulk on the P-phenyl groups of triphenylphosphine-derived ylides (e.g. using tri(ortho-toly)phosphine derived ylides) results in increased Z-selectivity. Thus, the factors that should exacerbate the need to relieve 1-3 steric interactions in a puckered cycloaddition interaction with C=O should be increased Z-selectivity, as this mechanism would predict.

(ii) Selectivity in reactions of semi-stabilised ylides

The Wittig reactions of semi-stabilised ylides are proposed to proceed by an irreversible [2 + 2] cycloaddition to give OPA, with C-C bonding being more advanced at the TS than P-O bonding. The initially formed OPA with oxygen in an axial position undergoes pseudorotation to place the ylidic carbon in an apical position in the phosphorus-centred trigonal bipyramid. This OPA then undergoes stereospecific and irreversible cycloreversion to alkene & phosphine oxide.\textsuperscript{99} The TS is thought to occur later along the reaction coordinate than in reactions of non-stabilised ylides. A consequence of this is that the phosporus substituents (spectator ligands, ylidic carbon and aldehyde oxygen) are in a tighter, more nearly trigonal bipyramidal arrangement about phosphorus than is the case in TSs of reactions of non-stabilised ylides (see Fig. 10). Also the shape of the sp\textsuperscript{2} hybridised phenyl or vinyl substituent at carbon-2 of the forming ring is postulated to result in less severe steric interactions between all of the substituents in the TS compared to the alkyl group in the same position in the TSs derived from non-stabilised ylides.\textsuperscript{93} It can be envisaged that the unsaturated group could orient itself to avoid steric interactions in a way that an alkyl group could not in such TSs. This may in turn have an effect on how the substituents on phosphorus are disposed. The upshot is that there is less of an energetic advantage to TS puckering, which in reactions of non-stabilised ylides nearly minimises both 1-3 and 1-2 steric interactions. The somewhat decreased importance of 1-3 (and presumably also of 2-3) interactions means that parallel (or nearly parallel) approach of the ylide and aldehyde can be competitive energetically with a TS involving the puckered approach of the reagents. In such a scenario the minimisation of 1-2 interactions dictates that a planar TS should be selective for trans-OPA (Fig. 10(b)). The cis-selective TS is likely to be puckered, and thus to be similar in appearance to the cis-selective TS in reactions of non-stabilised ylides.
The computational findings of Aggarwal, Harvey and coworkers\textsuperscript{[25]} are consistent with the operation of the above mechanism in the reaction of benzylidenediphenylphosphorane with benzaldehyde at the B3LYP/6-31G*(THF) level of theory. The cis-OPA selective cycloaddition pathway is marginally favoured over the trans-selective pathway. The barrier to OPA decomposition is lower than that to OPA formation. The activation barriers for both steps are lower than that for reversal of OPA to starting materials. Thus OPA formation is rate-determining (cf. reaction of ethylidenediphenylphosphorane with benzaldehyde indicating OPA formation is rate-determining and irreversible, and that the cycloaddition TS is indeed later (i.e. bond formation and rearrangement of substituent geometries about the reactive centres are more advanced) than for the corresponding reaction of a non-stabilised ylide.\textsuperscript{[42]} OPA formation was found to be exothermic, albeit not to the same extent as the in reaction of ethylidenediphenylphosphorane. The cis-selective TS shows only slight puckering, with a positive angle, while the energetically favoured trans-selective TS is also slightly puckered but in a negative sense. There is a possibility that the decrease in the puckering in the cis-selective TS is as a result of a reduction of the unfavourable electrostatic interaction in the “flatter” conformation between the C-O and ylide C-Ph bond dipoles see later).

The importance of 1-3 interactions is emphasised by the different selectivities observed in reactions of a given semi-stabilised ylide with primary and tertiary aldehydes respectively. The latter generally show much greater Z-selectivity than the former as a result of more pronounced 1-3 interactions. The mechanism described above for the Li-salt free Wittig reaction of semi-stabilised ylides is consistent with the following facts:

- Where OPAs can be observed spectroscopically (i.e. dibenzophosphole-derived OPAs, where the rate of decomposition is retarded sufficiently to make OPA decomposition rate determining), their formation has been shown to be under kinetic control, and stereospecific conversion to alkene has been proven.\textsuperscript{[90]}

- Stereospecific conversion of β-HPS (by deprotonation)\textsuperscript{[90,96]} or epoxide (by nucleophilic cleavage with phosphide, and methylation of the resulting β-oxidophosphine)\textsuperscript{[90]} to alkene via (presumed) betaine and OPA intermediates has been demonstrated.

- Attempted crossover experiments on Wittig reactions of semi-stabilised ylides gave no crossover product.\textsuperscript{[96,106]} although as previously mentioned, this may not be meaningful if the Wittig reaction is already complete by the time the crossover reactant is added.

- No intermediate is observed in Wittig reactions of semi-stabilised ylides (except for DBP ylides). Kinetic studies show the reaction to be overall second order, and to have a

In reactions of semi-stabilised ylides for which one or more of the P-phenyl groups are replaced by alkyl group(s), 1-3 interactions become less important, as the P-alkyl groups effectively free up space around phosphorus in the cycloaddition TS in a way that is not possible with three P-phenyl groups. This is thought to be more to do with the shape of the P-alkyl group and its effect on how the other substituents on phosphorus are oriented (so its effect on the shape of the “phosphonium” moiety as a whole) than its relative steric bulk per se. As a consequence of the above, the main steric interaction that destabilises the nearly planar trans-selective TS for triphenylphosphine-derived ylides is dramatically reduced for reactions of these ylides, and so it is energetically favoured over the cis-selective TS due to its much smaller 1-2 interactions. This results in much higher kinetic selectivity for the trans-OPA in the cycloaddition step, which is why high E-selectivity is observed for such semi-stabilised ylides. Computational results on the reaction of benzylidenedimethylphosphorane with benzaldehyde indicate OPA formation is rate-determining and irreversible, and that the cycloaddition TS is indeed later (i.e. bond formation and rearrangement of substituent geometries about the reactive centres are more advanced) than for the corresponding reaction of a non-stabilised ylide.\textsuperscript{[42]} OPA formation was found to be exothermic, albeit not to the same extent as the in reaction of ethylidenediphenylphosphorane. The cis-selective TS shows only slight puckering, with a positive angle, while the energetically favoured trans-selective TS is also slightly puckered but in a negative sense. There is a possibility that the decrease in the puckering in the cis-selective TS is as a result of a reduction of the unfavourable electrostatic interaction in the “flatter” conformation between the C-O and ylide C-Ph bond dipoles see later).

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- No intermediate is observed in Wittig reactions of semi-stabilised ylides (except for DBP ylides). Kinetic studies show the reaction to be overall second order, and to have a
positive \( \rho \) value for the carbonyl reagent.\(^{55,56,107}\) Negative activation volumes\(^{56}\) and entropies\(^{55}\) were also reported, which would be in line with a cycloaddition mechanism. Reactions with substituted benzaldehydes showed a kinetic isotope effect at the carbonyl carbon (\(^{14}\)C labelled in the study concerned), indicating that the bonding is changing at the carbonyl centre in the rate determining step.\(^ {107}\)

- No betaines derived from semi-stabilised ylides have been detected in the course of spectroscopic monitoring of Wittig reactions,\(^ {69,86}\) \( \beta \)-HPS deprotonation reactions or reactions involving epoxide ring opening and phosphorus quaternisation.\(^ {60,86}\)

- The reaction of \( P \)-chiral semi-stabilised ylide benzylidene-ethyldiethylphosphine oxide with benzaldehyde gave phosphine oxide in which there is retention of configuration at phosphorus. This indicates the operation of a syn-elimination of the phosphine oxide from a cyclic intermediate or TS with trigonal bipyramidal phosphorus.\(^ {79}\)

- The above points are consistent with an irreversible, rate-determining cycloaddition of ylide and aldehyde to give OPA where the aldehyde carbonyl is acting as an electrophilic centre, followed by an irreversible, stereospecific syn-cycloreversion of OPA to alkene and phosphine oxide.

- The experimentally observed selectivities are well accounted for by the proposed mechanism. It is consistent with the facts that the most \( Z \)-selective reactions of semi-stabilised ylides involve tertiary aldehydes or bulky triarylphosphine- or bulky aldehydes or dialkylphosphine- or methylidibenzophosphido-derived ylides, and that the most \( E \)-selective reactions of these ylides involve non-bulky aldehydes or dialkylphosphine- or trialkylphosphine or methylidibenzophosphido-derived ylides.

(iv) Selectivity in reactions of stabilised ylides

Stabilised ylides react through a relatively late TS, in which rehybridisation about phosphorus to trigonal bipyramidal and about the ylidal carbon and aldehyde carbonyl centres to tetrahedral is near completion. Vedejs’s original proposal contended that this meant the cycloaddition TS was constrained to being planar, and hence 1-2 interactions would generally be the dominating steric interactions, which would be consistent with the generally observed high \( E \)-selectivity in Wittig reactions of stabilised ylides.\(^ {60,89}\) However, some small inconsistencies did exist with this proposed mechanism; for example it could not adequately explain the high \( E \)-selectivity observed in reactions of \( \alpha \)-alkyl-\( \alpha \)-carbonyl substituted stabilised ylides. These inconsistencies were resolved by Aggarwal, Harvey and co-workers based on computational studies.\(^ {41,42}\) Notably, they modified the cycloaddition mechanism for stabilised ylides proposed by Vedejs and co-workers to take into account dipole-dipole interactions in the cycloaddition TS.

Calculations on the reaction of (methoxy carbonyl methylidene)-triphenylphosphorane with benzaldehyde at the B3LYP/6-31G* (THF) level of theory indicate rate-determining endothermic cycloaddition of ylide and aldehyde to give OPA, which undergoes facile pseudorotation to place the ylidal carbon in an axial position in the phosphorus-centred trigonal bipyramidal, and then cycloreverts to alkene and phosphine oxide, as predicted by the Vedejs mechanism. The cycloaddition TS is indeed found in their calculations to be relatively late (i.e. bond formation is quite advanced), and the

![Fig. 11](image-url) (a) Favoured trans-selective TS showing negative puckering angle, (b) Lowest energy cis-selective TS, (c) & (d) Higher energy cis-selective TSs.
The computational results on the reaction of (methoxy carbonyl)methylidene)trimethylphosphorane with benzaldehyde (B3LYP/6-31G+(THF) level of theory) indicates a similar shape for the cycloaddition TSs to the above reactions, and indeed a similar energetic advantage for the trans-selective TS over the cis, which is in keeping with the observed high E-selectivity in reactions of trialklyphosphine-derived ylides. However, it may be that in certain circumstances, if the phosphorus substituents are not as bulky as triphenylphosphine-derived ylides (e.g. for methyldiphenylphosphine-derived ylides), then a different cis-selective TS with an antiparallel orientation of reactant dipoles is possible (Fig. 1c) since placement of the large substituent in the pseudo-axial position is not discriminated against to the same extent by 1-3 interactions, and so somewhat lower E-selectivity observed.\(^4\)

Diminished E-selectivity, and even predominant Z-selectivity, has been observed in reactions of stabilised ylides in methanol. A reasonable explanation is that this results from solvent-induced decrease in the importance of the interaction of reactant dipoles in the cycloaddition TSs. In this scenario, the factors governing TS geometry may be quite similar to those in reactions of semi-stabilised ylides.

This mechanism, in taking account of reactant dipole-dipole interactions in the cycloaddition TS, allows a rationalisation of the consistently high E-selectivity observed in Wittig reactions of α-alkyl-α-carbonyl disubstituted stabilised ylides. In reactions of such ylides, 1-2 steric interactions alone cannot account for the observed bias in selectivity towards trans-OPA and E-alkene. The presence of the additional ylide α-substituent, however, does not alter the operation of the dipole-dipole interaction, so the trans-selective TS adopts a similar conformation to that in reactions of the mono-substituted ylide (see Fig. 12), and is similarly favoured energetically over the cis-selective TS for the same reasons as are present in reactions of the mono-substituted ylide.

Further experimental evidence for the operation of the cycloaddition mechanism in Wittig reactions of stabilised ylides has been presented recently in a publication quantifying the relative nucleophilicity of phosphonate carbanions (Wadsworth-Emmons reagents), phosphinoyl carbanions and stabilised phosphonium ylides towards substituted quinone methides (Michael acceptors), benzylidinium ions (carbocations) and benzaldehydes.\(^5\) These nucleophiles are obliged to react by a straightforward nucleophilic substitution with the former two electrophiles. The relative reactivity of all three types of phosphorus ylide towards the benzaldehydes was found to be systematically lower than towards the carbocations and Michael acceptors. Based on this, the authors concluded that the ylides all react with carbonyl compounds through an asynchronous concerted [2+2] cycloaddition. The \(\rho\) values determined in this study for the benzaldehydes in their reactions with each phosphorus nucleophile (2.9 for the phosphonium ylide, 3.4-3.6 for the phosphonate carbanions, and 2.7-2.8 for the phosphinoyl carbanions) are consistent with the operation of a cycloaddition mechanism with C-C bonding being more advanced than P-O bonding in the reactions of all three nucleophiles with carbonyl species. The value found for the reactions of the phosphonium ylide is consistent with literature precedent.\(^5\)

The mechanism described above for the Li-salt free Wittig reaction of stabilised ylides is consistent with the following facts:

- Stereoselective conversion of β-HPS (by deprotonation) to alkene via (presumed) betaine and OPA intermediates has been demonstrated. Thus OPAs are formed under kinetic control, and the stereoselectivity is decided in the C-C bond forming step.\(^6\)
- Reactions are second order (first order in each of ylide and carbonyl species), and alkene appears at the same rate as ylide is consumed.\(^5,\)\(^5\)
- \(\rho\) values for the carbonyl reactants have been found to be positive (so it is acting as an electrophile in the rate determining step)\(^5,\)\(^5\) and for ylides have been shown to be negative (so it is acting as a nucleophile),\(^5\) and their magnitudes are consistent with the operation of an asynchronous cycloaddition mechanism.
- The nucleophilicity of the stabilised ylide (methoxy carbonyl)trimethylphosphorane towards benzaldehydes has been found to be quantifiably different from its nucleophilicity towards carbocations and Michael acceptors, with which it must react by nucleophilic addition.\(^7\)
- The entropy of activation for reactions of stabilised ylides is large and negative.\(^5,\)\(^5\) There are no strong solvent effects in the reaction,\(^10\) and indeed it has been observed that the rate of the reaction of (fluorenylidene)ethyldiphenylphosphorane with \(p\)-nitrobenzaldehyde is slower in acetone or DMF than in benzene.\(^5\) Thus, the TS of the rate determining step is not polar and is highly ordered, characteristic of a cycloaddition TS.
- No OPA or betaine derived from a stabilised ylide has been detected during spectroscopic monitoring of Wittig reactions\(^6\) or β-HPS deprotonations.\(^6\)

5. Conclusions

It is our hope that this review will lead to the widespread recognition of the true course of the Li salt-free Wittig reaction, and that this will be reflected in the content of undergraduate textbooks. The fact that there is a clear distinction between the mechanisms that operate in the presence and absence of Li salts should be delineated, and in particular the fact that the Li-salt free mechanism is now definitively known while the Li-present mechanism is still effectively unknown should be clearly stated.

In summary: the first step of the Li-salt free Wittig reaction is an irreversible \([2 + 2]\) cycloaddition to give OPA. The stereochemistry of the product alkene is set during the formation of OPA. The varying shapes of the \([2 + 2]\) cycloaddition TSs lead to different diastereoselectivities in reactions of different
yldes. High selectivity for E-alkene is not necessarily indicative of the operation of thermodynamic control in any Li-salt free Wittig reaction; with only a small set of well-defined exceptions all such reactions are irreversible. High E-selectivity results from a kinetic preference for the formation of trans-OPA. The role of OPA pseudorotation in the Wittig reaction has yet to be settled, but it seems likely that OPA decomposition occurs in a single step that is comprised of both a Berry pseudorotation process and P-C and C-O bond breakage. What is certain about the OPA decomposition step is that it is stereospecific with respect to the carbon atoms of the newly-formed alkene.

The simplicity of this modern two-step interpretation of the Li-salt free Wittig reaction mechanism belies great complexity, in particular in the source of diastereoselectivity. It is nonetheless ironic in light of the many proposals made, words written, diagrams drawn and experiments done that it is, at its most basic level, perhaps the simplest possible means of turning ylide + carbonyl into alkene + phosphine oxide that has turned out to be the true mechanism.

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Notes and references

3 A. W. Johnson, Ylides and Imines of Phosphorus, 1993, Wiley; New York, chapters 8 and 9, pp. 221-305.
7 A. D. Abell, M. K. Edmonds in Modern Carboxyl Olefination; T. Takeda, Ed.; Wiley-VCH: Weinheim, Germany, 2004; Chapter 1
16 See reference 3, chapter 8, pp.221-273.
25 R. Bera, G. Dhananjaya, S. N. Singh, R. Kumar, K. Mukkanti, M. Pal, Tetrahedron 2009, 65, 1300-1305: describes the use of microwave conditions and is a good leading reference to other variations