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The role of Bronstead Acidity in poisoning the SCR-Urea reaction over FeZSM-5 catalysts.

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Abstract:
Two series of FeZSM-5 catalysts prepared from Na⁺ and NH₄⁺ ZSM-5 precursors are studied in the selective reduction of NOₓ using NH₃ and urea as reducing agents. All Fe-containing catalysts are active for NOₓ reduction in the SCR-NH₃ reaction with ex-NH₄⁺ catalysts being more active than ex-Na⁺ materials and the activity depending (to a minor extent within each series of catalysts) upon [Fe]. Catalysts with Bronstead acid sites also show a small transient deNOₓ activity at low temperatures. All catalysts are less active for the SCR-Urea reaction but the ex-Na⁺ catalysts retain far more deNOₓ activity than the ex-NH₄⁺ materials. NH₃ TPD shows that strongly binding Bronstead acid sites are present on the ex-NH₄⁺ materials and H⁺-treated parent zeolites while urea TPD shows that the mode of decomposition of urea differs as a function of initial zeolite counter-ion. Urea TPSR shows that the reaction between adsorbed urea and gaseous NO/O₂ is related to [Fe]. It is proposed that the decreased activity of the ex-NH₄⁺ catalysts in the SCR-Urea reaction is due to a less favourable mode of decomposition over these catalysts. Furthermore it is suggested that the Bronstead acidity plays some part in this less favoured decomposition.

Keywords: FeZSM-5, deNOₓ, Urea, NH₃, TPD.
Introduction:

Fe-ZSM5 materials have been used in several selective oxidation reactions, e.g. where N$_2$O is used to oxidise benzene to phenol\textsuperscript{1} and in the oxidative dehydrogenation of propane to propene\textsuperscript{2}. This material has also been studied in the Selective Catalytic Reduction of NO$_x$ using both hydrocarbons\textsuperscript{3-5} and NH$_3$ as reducing agents\textsuperscript{6-8}.

Selective Catalytic Reduction involves the reduction of NO$_x$ in the presence of a large excess of O$_2$. This reaction is important because of the detrimental effects of NO$_x$ on the atmosphere (photochemical smog) the aqueous environment (acid rain) and human health (lung ailments). There are a large number of situations where NO$_x$ exists in the presence of a large excess of O$_2$. These include the exhausts of stationary power sources, diesel and lean-burn gasoline exhausts where NO$_x$ is formed at high temperatures from the NO + O$_2$ reaction via the Zeldovich mechanism, and the exhausts of FCC regeneration reactors where catalyst bonded coke is oxidised at high temperatures from the surface in order to regenerate active cracking catalysts. Any N-containing organic species in the coke will, under these conditions, be oxidised to NO$_x$\textsuperscript{9}.

The use of NH$_3$ as a selective reducing agent is feasible for large-scale power generation plants or FCC recycling units since these are, by definition, stationary and, once operational, they produce a constant amount of NO$_x$. Thus, NH$_3$ containment and dosage issues are much lessened. On board a vehicle it would not be possible to transport cylinders of NH$_3$ for use in exhaust after-treatment and for this reason many researchers are beginning to look at solutions of (NH$_2$)$_2$CO (urea) as selective reduction agents\textsuperscript{10-12}. In the presence of H$_2$O, urea should decompose according to equation 1 to form 2 NH$_3$ and one CO$_2$ molecule\textsuperscript{13}.
\[(\text{NH}_2)_2\text{CO} + \text{H}_2\text{O} \rightarrow 2\text{NH}_3 + \text{CO}_2\]  

This reaction is reported to take place in two steps

\[(\text{NH}_2)_2\text{CO} \rightarrow \text{NH}_3 + \text{HNCO}\]  

\[\text{HNCO} + \text{H}_2\text{O} \rightarrow \text{NH}_3 + \text{CO}_2.\]

Solutions of urea are non-toxic, non corrosive and odour-free and can be relatively concentrated and therefore may well provide a usable on-board reductant for the removal of NO\textsubscript{x} from diesel engines and lean-burn-gasoline vehicles.

Recently Larsen et al. have studied the reactions of urea and NO over several zeolitic-type materials, including HZSM and FeZSM\textsuperscript{14} using MAS NMR. In their studies these workers noted the production of NH\textsubscript{3} and CO\textsubscript{2} following the interaction between urea and FeZSM and H-ZSM and the formation of NH\textsubscript{3}, CO\textsubscript{2} and N\textsubscript{2} over the same materials when the surface was dosed with both NO and urea. Production of N\textsubscript{2} was not seen over silicalite and these authors suggest that Bronstead acidity is important in promoting the desired reaction. Resonances due to the formation of melamine were observed after heat treatment of the silicalite in urea and NO/O\textsubscript{2}\textsuperscript{14}.

In this work we have compared the efficiency of FeZSM-5 catalysts for the reduction of NO\textsubscript{x} using both NH\textsubscript{3} and equivalent amounts of (NH\textsubscript{2})\textsubscript{2}CO. The catalysts, and their interactions with NH\textsubscript{3} and urea, are characterised using NH\textsubscript{3} Temperature Programmed Desorption and Urea Temperature Programmed Decomposition and Temperature Programmed Surface Reaction. Our results suggest that materials with strong Bronstead acidity are less active in the SCR-Urea reaction.

**Experimental:**
Catalyst Preparation: Aliquots of two samples of commercially available ZSM-5 zeolites (AlsiPenta) were used as parent zeolites. These differ according to their initial parent counter ion, e.g. in one the parent counter-ion is NH$_4^+$ and in the second the parent counter-ion is Na$. In each case the SiO$_2$/Al$_2$O$_3$ ratio is the same, i.e. 27 and thus the ion-exchange capacity of both zeolites is the same. Fe was introduced to the parent zeolites using standard ion-exchange techniques, i.e. the samples were shaken for 2 h in solutions of FeSO$_4$ of varied volume and concentration, then filtered, washed in deionised water to remove any residual counter ion (SO$_4^{2-}$), dried at 120 °C for 2 h to remove H$_2$O, and calcined at 500 °C for 2 h. The powder was then crushed and sieved and the fraction of catalyst between 212 and 600 µm were used in the subsequent analyses. Two “blank” catalysts were prepared by ion-exchanging the parent zeolites in H$_2$SO$_4$ (at equivalent SO$_4^{2-}$ concentrations to those used in the FeZSM-5 preparations), washing, drying and calcining as before. The [Fe] within the zeolites following preparation was determined by Atomic Absorption spectroscopy following acid digestion.

One unusual feature, relating to a difference between the ex-Na$^+$ and ex-NH$_4^+$ catalysts, was noted during these preparations. In situations when the exact same protocol was used to prepare Fe-loaded catalysts the same mass of parent zeolite was shaken in an Fe solution of the same volume and concentration for the same period of time. Then the samples were washed in the same volume of hot distilled water, dried and calcined. Invariably catalysts prepared from parent zeolite with the NH$_4^+$ counter ion contained less Fe than those prepared from the parent zeolite with the Na$^+$ counter ion. As mentioned the SiO$_2$/Al$_2$O$_3$ ratio (and thus the ion-exchange capacity) of both parent zeolites is the same and thus some other feature must account for the differing concentrations in the final catalysts. One possible explanation of this difference lies in
the fact that the pH of the exchanging solution changes as NH₃ is exchanged out of the samples while it remains essentially invariant during the Na⁺ exchange. This changing pH might in some way influence the nature of the incoming Fe ions (perhaps causing precipitation of FeOₓ crystallites) and preventing further exchange. Similar effects have been noted when CuZSM-5 catalysts are prepared from these different parent materials.

**Activity measurements:** The catalyst (25 mg) was held in a tubular quartz reactor using plugs of quartz wool. The reactants were blended using electronic mass flow controllers from cylinders of 1% NO and 1% NH₃ in He (BOC Special Gases), O₂ and He (BOC) to give a reaction mixture of [NO] = 1000 ppm, [NH₃] = 1000 ppm and [O₂] = 13% in a total flow of 100 ml/min. When solutions of urea were used as a reductant the liquid was introduced into a heated zone before the reactor from a calibrated syringe driver. Aqueous solutions of 5% urea were used at liquid flow rates which gave 1000 ppm urea in the final reaction mixture. The reactions (SCR-NH₃ and SCR-Urea) were studied under temperature-programmed conditions with a ramp rate of 10 °C/min (30 –500 °C for the SCR-NH₃ reaction and 120 – 500 °C for the SCR-Urea reaction. The NOx concentration in the exit gas was continuously monitored, following suitable dilution, using a Monitor Labs Inc. Nitrogen Oxides Analyser (Model 8840) connected to a PC.

**NH₃-TPD:** In these experiments the catalyst (50 mg) was dosed with NH₃ (2860 ppm) at 110 °C for 30 minutes. The NH₃ was then removed from the stream and the catalyst cooled to 50 °C in a flow of He (70 ml / min). The dosed catalyst was held at this temperature for 20 minutes in the He flow (until the NH₃ signal returned to the baseline) and then the temperature ramped from 50 to 550 °C at a ramp rate of 10 °C / min. Desorbed NH₃ is continually monitored as a function of temperature.
using a Prolab mass spectrometer running in the multiple ion monitoring mode. The data is treated to remove interference from a signal for H$_2$O$^{16}$ and then converted into ppm-NH$_3$ for presentation.

*Temperature Programmed Urea Decomposition:* The interaction of the catalysts with urea was probed using Temperature Programmed Urea Decomposition$^{12}$. In the case of a urea TPD the catalyst was repeatedly wet with an aqueous solution of (5%) urea a sufficient number of times to dose the 50 mg catalyst sample with 50 µmol of urea. The sample was then dried at 50 °C and placed in the reactor in a flow of 70 ml/min of He at a temperature of 50 °C. The temperature was ramped from 50 to 750 °C at a rate of 20 °C / min. The effluent gas was passed through a condenser at 0 °C in order to remove gas phase urea (to prevent its condensation within the capillary or the mass spectrometer chamber). Thus, only the products of the interaction between urea and the catalyst were monitored during decomposition.

The effluent gas was monitored continuously by mass spectrometry. Masses at 15, 16, 17, 18 (H$_2$O, NH$_3$ and various fragments of each), 28 (N$_2$, CO), 30 (NO), 32 (O$_2$) and 44 (CO$_2$, N$_2$O) were monitored as a function of temperature. The data were then corrected for overlapping masses (*e.g.* the contribution of the H$_2$O fragment at 17 to the NH$_3$ signal was removed)$^{16}$. Following such corrections the data were converted into ppm readings using conversion factors determined by calibration. No calibrations were carried out for the HNCO intermediate so any data presented showing ppm or µmol HNCO are an approximation. In any case HNCO is never a major product of either the TPD or the TPSR experiments.

In order to discriminate between N$_2$ and CO profiles (and N$_2$O and CO$_2$ profiles) these experiments were repeated using a dose of $^{15}$N labelled urea (99
atom% $^{15}\text{N}$ Aldrich). The plots presented here are combinations of those observed using the labelled and unlabelled urea dosages.

**Urea TPSR:** As a final characterisation the temperature programmed surface reaction of $(\text{NO} + \text{O}_2)_{(g)}$ with adsorbed urea was studied. In these experiments 10 mg of the catalyst was dosed with 50 $\mu$mol urea and then ramped from 50 °C to 750 °C in a flow of NO + O$_2$ (1000 ppm NO + 10% O$_2$) at 20 °C min$^{-1}$.

**Results and Discussion:**

Figure 1 shows the activity of all the catalysts in the temperature programmed SCR-NH$_3$ reaction. Both parent zeolites show a transient low temperature activity and generally comparable high temperature activity (with the ex-Na$^+$ catalyst being slightly more active than the ex-NH$_4^+$ at 500 °C). In the absence of Fe ions the treatment of both in dilute H$_2$SO$_4$ solutions generates catalysts with similar behaviours. This observation is again made later in the characterisation experiments and suggests that the acid treatment generated similar sites within both parents (through removing NH$_4^+$ and Na$^+$ and replacing these ions with H$^+$). Specifically the transient lower temperature activity is associated with Bronstead acid sites.

Within each zeolite family those catalysts containing Fe are more active than the proton exchanged materials. Also in both cases the % Fe exchanged within the zeolite does not dramatically affect the catalyst activity, i.e. in the case of the ex-NH$_4^+$ catalysts the 0.6% Fe catalyst is slightly more active than the 0.3% analogue while in the case of the ex-Na$^+$ materials the 1% Fe catalysts is slightly more active than the 0.7% analogue. The ex-NH$_4^+$ materials (even though these have lower Fe concentrations) were more active than the ex-Na$^+$ catalysts, reaching 10% conversion at 240 °C (compared to 280 °C for the ex Na$^+$) and 50% conversion at 365 °C.
(compared to 395 °C for the ex-Na⁺). They also show a higher maximum NOx conversion ~100% at 450 °C compared with ~90% at 495 °C.

These ex-NH₄⁺ Fe-containing materials also show a lower temperature for turnover of activity (where the selectivity of the reaction changes and NH₃ begins to be combusted (470 compared to 500). They also show a transient low temperature activity between 250 and 280 which is absent in the ex-Na⁺ materials. Strong Bronstead acid sites (as measured by NH₃ TPD) are present in all the samples which show this transient activity.

Similar transient low temperature activity peaks have been seen previously over high SiO₂–containing CuZSM-5 catalysts¹⁵ and (albeit at lower temperatures) in the SCR-NH₃ reaction over FeZSM-5¹⁷.

The temperature programmed SCR-Urea profiles are shown in Figure 2. There is a dramatic difference between the activities of both families of catalysts in the SCR-Urea reaction. Fe-containing materials prepared from Na-ZSM-5 are far more active than those prepared from NH₄⁺-ZSM-5.

Again both parent zeolites show similar activity (including a transient increase in activity at 350-400 °C). In the case of the Fe-containing Na⁺-derived zeolites the deNOx activity profile is shifted upwards in temperature by ~50 °C relative to activity in the SCR-NH₃ reaction but in the case of the Fe-containing ex-NH₄⁺ materials the activity is almost totally quenched (notwithstanding the transient activity over the 0.6% Fe-ZSM catalyst at ~ 400 °C).

The reasons for these differences between activity in the SCR-NH₃ reaction and activity in the SCR-Urea reaction must originate in the manner in which urea interacts with the different catalysts since the presence of H₂O in the SCR-NH₃ reaction mixture (the only other difference between the reactants impinging on the
catalysts in both cases) has very little effect on activity, i.e. a slight promotion of activity over ex-Na\(^+\) catalysts and a slight poisoning effect over ex-NH\(_4\)\(^+\) materials (results not shown)\(^{15}\). Similar results have been noted when comparing the SCR-NH\(_3\) and SCR-Urea reactions over supported copper oxide catalysts\(^{18}\).

In an attempt to gain an insight into the NH\(_3\) and Urea handling characteristics of both series of catalysts a set of NH\(_3\) and Urea Temperature Programmed Decomposition studies were carried out. The results of the NH\(_3\) desorption experiments are presented in Figure 3. The shapes of the profiles are in good agreement with those recorded previously\(^{6,7}\). Both families of catalysts show desorption of NH\(_3\) at several temperatures suggesting that there are different types of NH\(_3\) coordination sites on the materials. No NH\(_3\) oxidation products are detected during the temperature ramp.

The two parent zeolites both show three types of desorption with features at 140 °C, 220 °C and 445 °C. The latter feature is far stronger in the case of the ex-NH\(_4\)\(^+\) material than from the ex-Na\(^+\) catalyst suggesting that it relates to Bronstead bound NH\(_3\) and that the concentration of sites available to form this type of NH\(_3\)\(_{\text{ads}}\) is far higher over this catalyst than over the non-acidic starting material. Thus we can say that even though the treatment in H\(_2\)SO\(_4\) did confer some Bronstead acidity on the Na\(^+\)-ZSM-5 system (generation of a high temperature desorption peak), not all the sites able to form Bronstead sites in this material do so.

The high temperature peaks (which derive from desorption of a Bronstead acid bound NH\(_3\) species\(^{19}\)) are not present in the FeZSM-5 samples prepared from NaZSM-5 while they are present in the decomposition profiles seen from the ex-NH\(_4\)\(^+\) catalysts. The ex-NH\(_4\)\(^+\) Fe-containing samples also show two less well defined peaks at lower temperatures (140 °C and 200 °C) that are also present in the parent zeolite.
It is not possible to deconvolute the contributions here from any Fe-NH$_3$ surface complexes.

For the ex-NH$_4^+$ catalysts the overall acidity (as measured by the total amount of NH$_3$ desorption) decreases as Fe loading is increased but this is not the case for the ex-Na$^+$ materials (see Table 1). There is also an obvious and expected decrease in the concentration of sites which form the NH$_3^{ads}$ which is stable at high temperature - as protons are exchanged by Fe ions Bronstead acidity decreases. There are also two lower temperature peaks present in the TPD profiles seen from the ex Na-ZSM-5 catalysts. The temperature distribution of these peaks is wider than that seen over the NH$_4^+$-derived catalysts with peaks at 160 °C and 280 °C.

These differing desorption profiles show that there are differences in the interactions between NH$_3$ and the two types of Fe-containing materials. The most striking difference is the presence of a high temperature peak relating to the decomposition of a Bronstead-acid bound species in the ex-NH$_4^+$ catalysts which is absent in the ex-Na$^+$ materials. The presence (or absence) of this type of site obviously has very little detrimental impact on the activity of the catalyst in the SCR-NH$_3$ reaction; indeed it promotes this reaction and contributes to the transient lower temperature activity (Figure 1). However it seems that the presence of this type of site, possibly through unfavourable interactions with the basic urea molecules, causes the decreased deNOx activity of the catalysts when urea is used as a reductant.

In order to clarify the interactions between the two families of catalysts and urea a series of urea Temperature Programmed Decompositions were carried out. For clarity only representative plots are shown, i.e. those relating to the 0.6% ex-NH$_4^+$ catalyst (figure 4) and 0.7% ex-Na$^+$ catalyst (figure 5). However, the other ex-NH$_4^+$ and ex-Na$^+$ Fe-containing catalysts behave similarly to these two examples. The two
parent zeolites also behave in a similar manner to one another and the profiles shown in figure 6 shows how urea interacts with the H-exchanged zeolite. A large water evolution was found to peak at ~ 150 °C presumably due to physisorbed H₂O within the pores of the zeolite. There was also a smaller production of H₂O peaking at ~450 °C which can be ascribed to zeolitic dehydroxylation. These peaks have been found over all catalysts and are not shown for clarity. Decomposition products such as N₂O and CO are not found during these experiments (as probed using the labelled urea in order to discriminate between N₂/CO and N₂O/CO₂) and there was also negligible production of NO. Thus the main peaks of interest were CO₂, N₂, NH₃ and HNCO.

Figure 4 shows the main decomposition products seen during the temperature ramp over the 0.6% ex-NH₄⁺ Fe ZSM-5 catalyst. The production of CO₂ seems to be a complex process with four bands visible during the temperature ramp. The first band peaks at 200 °C and the last at 450 °C. N₂ production commenced at ~ 400 °C and rose to a steady state which seemed to continue for the remainder of the experiment. The production of NH₃ begins at ~ 300 °C and rises to a peak at ~ 490 °C. This temperature of desorption (suggested to be from Bronstead sites) differs from that seen during the NH₃-TPD experiments (450 °C) but this may be an artefact of the different ramp rates used during these experiments (20 °C min⁻¹) and that used during NH₃ TPD (10 °C min⁻¹). This discrepancy might also suggest the decomposition of some NH₃ precursor at this temperature. However, no other gas phase species are detected in parallel with this NH₃ peak, suggesting that NH₃_ads → NH₃(g) is the process being observed. The production of a relatively small amount of HNCO is noted between 200 °C and 250 °C (note the designation of this peak as representing ppm is an approximation).
If the adsorbed urea decomposed as suggested above (equations 1-3) it seems that any NH₃ formed during the initial decomposition is trapped on the zeolite and not released until a higher temperature. If the CO₂ is released immediately following initial urea decomposition (as seems likely since there are no surface basic sites available to trap it) then the profiles suggest that there is more than one type of catalytic site involved in the urea decomposition (leading to a four band profile for CO₂ evolution).

The peaks seen from the urea decomposition from the parent ex-NH₄⁺ zeolite (Figure 6) show very similar trends regarding NH₃ evolution (including commencement of NH₃ production and T_{max}), N₂ production, CO₂ production (the same four bands observed) and HNCO production. This suggests that it is the zeolite itself is also able to catalyse the decomposition under these conditions.

Different profiles are seen in the case of the ex-Na⁺ Fe-containing catalysts as represented by the 0.7% Fe ZSM catalyst profiles in Figure 5. There is no production of HNCO seen during the ramp. The production of CO₂ takes place in only two (rather than four) stages (peaking at 220 °C and 350 °C). NH₃ formation and desorption is at a far lower temperature (peaking at 300 °C). The latter result can be predicted from the TPD analysis where it was seen that the thermal stability of adsorbed NH₃ was lower on ex-Na⁺ Fe-containing zeolites (since they lack the Bronsted acidity inherent in the ex-NH₄⁺ materials). However, the lack of a peak due to HNCO is a significant difference between the profiles seen from the ex-NH₄⁺ and the ex-Na⁺ catalysts.

Thus these urea TPD experiments show that the decomposition of urea takes place in different manners over the two families of zeolite. These decompositions are significantly different to those seen over oxide supported CuO catalysts where most of
the NH₃ and CO₂ decomposition products are removed from the catalyst surface at a temperature of ~250 °C. The high temperature peaks seen during these reactions suggest that the interaction between urea and the zeolitic catalysts is far greater than the interactions between urea and the oxide supported CuO catalysts. This interaction presumably takes place between the acidic sites of the zeolites that are not present on the oxide catalysts. However, not all the urea that is dosed onto the catalyst remains anchored to the surface and available for decomposition. Previously we have noted that urea that is not co-ordinated to the catalyst surface simply sublimes from the reactor under the conditions of the experiment.

Table 2 shows the amounts of CO₂, NH₃, N₂ and HNCO produced during the temperature ramp. Note the figures given for HNCO are an approximation since the mass spectrometer was not calibrated for HNCO and the calibration factor for NO, is used to generate these figures. The error that this introduces is not large since the actual production of HNCO is relatively minor. The calibration factor for ¹⁵N₂ is assumed to be the same as that for ¹⁴N₂.

The catalysts were dosed with 50 µmol of urea sufficient, when reacted with water, to produce 50 µmol CO₂ and 100 µmol NH₃. It can be seen from the table that in no case are such amounts produced. The difference between the values expected and those observed can be accounted for by the Urea that sublimed from the reactor without decomposing as well as any fragments that remain on the surface following the temperature ramp (possibly through melamine formation).

It would be expected following a “clean” decomposition that the amounts of NH₃ produced should be roughly twice those of CO₂. While in most cases there is more NH₃ than CO₂ produced the ratios are not as would be expected (even when allowing for the production of N₂ and small amounts of HNCO). The 1% ex-Na⁺
FeZSM-5 catalyst decomposes the most urea but there does not seem to be any correlation between the amounts of urea which decompose from the catalysts and the [Fe] or the nature of the initial zeolite.

One possible explanation for this discrepancy is the formation of melamine passivating layers on the surface as the temperature is raised. This has previously been regarded as one of the modes of deactivation of SCR-urea catalysts and the layer is proposed to form through reaction between isocyanic acid (HNCO) and NH$_3$ (possibly through the formation of cyanuric acid – the cyclic trimer of HNCO).

To further analyse how these different decomposition pathways affect the SCR-NH$_3$ and SCR-urea activities of the catalysts Temperature Programmed Surface Reaction (TPSR) experiments were carried out. In these experiments the catalyst surface was dosed with urea and then ramped in a flow of NO and O$_2$ (1000 ppm NO, 10% O$_2$). It is not possible to directly compare the TPD results above with the TPSR results here since the urea is dosed onto different masses of catalyst. This means that a larger proportion of the urea in this experiment is not chemisorbed and as stated urea that is not chemisorbed simply sublimes from the reactor under these conditions. Qualitative comparisons are possible. The profiles recorded for the ex-NH$_4^+$ Fe ZSM (0.6% Fe) are shown in figure 7(a) and 7(b).

Figure 7(a) shows the profiles relating to N$_2$, CO$_2$ and NO. The CO$_2$ formation peaks are less well defined than in the case of the Urea TPD seen in Figure 4 but there are still more than one such peak suggesting more than one type of urea decomposition site. Again it seems that any NH$_3$ formed during the decomposition is held on the catalyst rather than being immediately released. The N$_2$ and NO profiles mirror one another with main peaks and lower temperature shoulders at (375 °C and ~260 °C in the case of the sample shown). The absolute amounts of NO removed
form the stream and N₂ and CO₂ formed also mirror one another with roughly 7 µmol of NO being removed and between 6 and 7 µmol of CO₂ and N₂ being formed. The low temperature shoulder is present in the case of all ex-NH₄⁺ Fe-containing materials and both “parent” zeolites but is absent in the case of the ex-Na⁺ Fe-containing materials suggesting it is related to the Bronsead acidity of the sample (specifically from the interaction between NO and the species which led to HNCO formation in the Urea decomposition experiments above).

The formation of NH₃ and HNCO in the TPSR experiments is much decreased (roughly two orders of magnitude) relative to those seen in the Urea decompositions (compare figure 7(b) with figure 4).

It seems that instead of forming NH₃(g) and HNCO(g) as was the case in the Urea TPD experiments these decomposition products go on to react with NO(g) to form N₂ leading to a peak in N₂ production from the NH₃ + NO(g) reaction at relatively high temperature (with the low temperature peak present over the materials that form HNCO being due to the HNCO + NO(g) reaction).

Thus, even though the ex-NH₄⁺ Fe-containing materials are unable to catalyse the steady state SCR-urea reaction that they are able to react gaseous NO with an amount of adsorbed urea in a relatively efficient manner. This suggests that a steady state amount of urea (or urea decomposition products) are required to form the proposed deactivating melamine layer.

The formation of N₂ and CO₂ and removal of NO is also seen over the ex-Na⁺ materials (figure 8). However the low temperature “bump” in NO removal and N₂ formation and is not seen (since HNCO is not formed over these materials) and the temperature of maximum reaction is slightly lower in this case (350 °C as against 380 °C). The amount of NH₃ being produced is also significantly higher than that seen in
the same experiment over the ex-NH$_4^+$ catalyst (while being significantly less than what was produced during urea TPD over the same material (figure 5)). It is possible that the desorption of NH$_3$ (rather than its reaction to N$_2$) takes place because over these materials the NH$_3$ is not held on the strongly interacting Bronstead sites (these not being present) and thus can desorb before sufficient energy is available to cause the NO$_{(g)}$ + NH$_{3ads}$ reaction to take place. In all other cases (results not shown) the temperature of maximum reaction decreases as the concentration of Fe increases.

**Conclusions:**

Fe ZSM-5 catalysts prepared from both Na$^+$ and NH$_4^+$ ZSM-5 materials are active catalysts in the SCR-NH$_3$ reaction. The activity of the catalysts depends on the initial starting material (with the ex-NH$_4^+$ catalysts being slightly more active than the ex-Na$^+$ materials) and, to a relatively minor extent, upon the concentration of Fe (with activity being directly proportional to [Fe]. When urea was used as a reducing agent the ex-NH$_4^+$ catalysts lost most of their deNOx activity while the activity of the ex-Na$^+$ materials was only partially suppressed.

NH$_3$ TPD has shown the presence of Bronstead acid sites within the ex-NH$_4^+$ zeolite (and also shows that an acid treatment can generate Bronstead sites within Na-ZSM-5). Temperature programmed urea decomposition showed that those catalysts which contain Bronstead acid sites also generate HNCO during urea decomposition. HNCO and NH$_3$ are intermediates in formation of melamine$^{21}$ and it is proposed that layers of this may be formed over these catalysts resulting in decreased activity in the SCR-urea reaction.

Temperature programmed surface reaction shows that it is possible under transient conditions to react these HNCO (and NH$_3$) adsorbed species with gas phase
NO/O₂ to form N₂ over all the catalysts with the temperature of maximum reaction being inversely proportional to the [Fe]. This suggests that under these transient conditions experiments do not mimic the conditions seen under steady state reaction where the proposed melamine deactivating layers form.

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References:
Figure 1: Temperature Programmed SCR NH$_3$ activity of the various catalysts (10 mg) (a) ex-NH$_4^+$ ZSM-5 catalysts, 0% Fe (■), 0.3% Fe (●), 0.6%Fe (▲), and (b) ex-Na$^+$ ZSM-5 catalysts 0% Fe (□), 0.7%Fe (○), 1.0% Fe (△).
Figure 2: Temperature Programmed SCR-Urea activity of the various catalysts

(10mg) (a) ex-NH$_4^+$ ZSM-5 catalysts, 0% Fe (■), 0.3% Fe (●) and 0.6% Fe (▲) and

(b) ex-Na$^+$ ZSM-5 catalysts, 0% Fe (□), 0.7% Fe (○) and 1.0% Fe (△).
Figure 3: Temperature Programmed NH$_3$ desorption from the various catalysts ex-NH$_4^+$ ZSM-5 catalysts, 0% Fe (■), 0.3% Fe (●) and 0.6% Fe (▲) and (displaced upwards) ex-Na$^+$ ZSM-5 catalysts, 0% Fe (□), 0.7% Fe (○) and 1.0% Fe (△).
<table>
<thead>
<tr>
<th>% Fe</th>
<th>total NH$_3$ desorbed / µmol g$^{-1}$</th>
<th>NH$_3$ desorbed at ~400 °C / µmol g$^{-1}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>ex-NH$_4^+$</td>
<td></td>
<td></td>
</tr>
<tr>
<td>0</td>
<td>1923</td>
<td>890</td>
</tr>
<tr>
<td>0.3</td>
<td>1790</td>
<td>829</td>
</tr>
<tr>
<td>0.6</td>
<td>1557</td>
<td>696</td>
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<tr>
<td>ex-Na$^+$</td>
<td></td>
<td></td>
</tr>
<tr>
<td>0</td>
<td>1544</td>
<td>584</td>
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<tr>
<td>0.7</td>
<td>1258</td>
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</tr>
<tr>
<td>1</td>
<td>1745</td>
<td>0</td>
</tr>
</tbody>
</table>

Table 1 showing the overall amounts of NH$_3$ desorbed (µmol g$^{-1}$) for each catalyst.

Also showing amounts desorbed in the high temperature peak associated with Bronsted acidity.
Figure 4: Main peaks of interest following temperature programmed decomposition of urea over ex-NH$_4^+$ FeZSM-5 (0.6% Fe). NH$_3$ (圜), N$_2$ (from an experiment using labelled urea) (□), CO$_2$ (◇) and HNCO (△) (estimate) (similar profiles were seen from the 0.3% ex-NH$_4^+$ Fe material).
Figure 5: Main peaks of interest following temperature programmed decomposition of urea over ex-Na\(^+\) FeZSM-5 (0.7% Fe). \(\text{NH}_3\) (\(\bigtriangleup\)), \(\text{N}_2\) (from an experiment using labelled urea) (\(\Box\)), \(\text{CO}_2\) (\(\bigtriangledown\)) and \(\text{HNCO}\) (\(\bigtriangleup\)) (estimate). Similar profiles were seen from the 1.0% ex-Na\(^+\) Fe-containing material).
Figure 6: Main peaks of interest following temperature programmed decomposition of urea over ex-NH$_4^+$ H ZSM-5. NH$_3$ (○), N$_2$ (from an experiment using labelled urea) (□), CO$_2$ (∙) and HNCO (∆) – (estimate) (similar profiles were seen from the ex-Na$^+$ H ZSM-5 material).
<table>
<thead>
<tr>
<th></th>
<th>CO₂</th>
<th>NH₃</th>
<th>N₂ *</th>
<th>HNCO **</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ex-NH₄⁺, 0.3%Fe</td>
<td>23.9</td>
<td>33.0</td>
<td>2.3</td>
<td>1.2</td>
</tr>
<tr>
<td>Ex-NH₄⁺, 0.6%Fe</td>
<td>13.1</td>
<td>22.1</td>
<td>3.0</td>
<td>1.1</td>
</tr>
<tr>
<td>Ex-NH₄⁺, Parent</td>
<td>16.2</td>
<td>24.1</td>
<td>2.0</td>
<td>1.6</td>
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<tr>
<td>Ex-Na⁺ 0.7%Fe</td>
<td>22.9</td>
<td>29.1</td>
<td>3.4</td>
<td>0.0</td>
</tr>
<tr>
<td>Ex-Na⁺ 1%Fe</td>
<td>33.6</td>
<td>33.9</td>
<td>3.2</td>
<td>0.0</td>
</tr>
<tr>
<td>Ex-Na⁺  Parent</td>
<td>19.8</td>
<td>28.1</td>
<td>1.1</td>
<td>0.5</td>
</tr>
</tbody>
</table>

Table 2: Amounts of NH₃, CO₂, N₂ and HNCO seen in the gas phase during Temperature Programmed Decomposition of 50 µmol Urea. *¹⁵N₂ from an experiment using labelled urea, ** approximation since accurate calibration factors for HNCO are not known.
Figure 7: Main peaks of interest following Temperature Programmed Surface Reaction of Urea_{ads} + NO/O_{2(g)} over ex-NH_{4}^{+} Fe ZSM-5 (0.6% Fe). (a) N_{2} (□), CO_{2} (◊) and NO (◆) (b) NH_{3} (○) and HNCO (△) (estimate).
Figure 8: Main peaks of interest following Temperature Programmed Surface Reaction of Urea$_{ads}$ + NO/O$_2$ (g) over ex-Na$^+$ Fe ZSM-5 (0.7% Fe). (a) N$_2$ (□), CO$_2$ (◊) and NO (◆) and NH$_3$ (○).