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The selective transfer of a methyl or alkyl group to an unactivated carbon center is of prevalent synthetic and biochemical interest.[1] In biological systems, methylation typically involves transferring a CH₃ group from a sulfide carrier such as S-adenosylmethionine (AdoMet) as mild methylating agent,[2] using a cobalamin cofactor as CH₃-donor.[3] While most AdoMet-type transfers deliver a methyl group, recent work showed that certain enzymes rather utilize a methylene unit (CH₂) for substrate alkylation.[4]

![Image of iridium complex]

**Scheme 1.** Generic methyl transfer, E' = S, N for mild transfer, E' = O, halide, for harsh transfer.

Synthetic mimics of methyltransferases have been developed for the transfer of a methyl group from sulfonyl or iminium salts to a heteroatom receptor (E = N, O, P, S; Scheme 1).[5] Non-enzymatic mimicking of alkyl group transfer from sulfur or nitrogen to carbon and formation of a new C–C bond, as observed for example in DNA methylation,[6] is very rare.[7] A key challenge is the E'–C₉₅ bond cleavage from the carrier system (E' = S, N), which is required to activate the transferable group.[8] Notably, selective C–N bond cleavage has been observed using an imidazolium-type source for alkyl group release under mild conditions in N-heterocyclic carbene (NHC) ruthenium complexes,[9] though no controlled transfer to a substrate was noticed. Here we report on an iridium complex that facilitates the selective transfer of a methylene group from a pyridinium fragment to an aryl unit with concomitant activation of a transferable group. This process involves C–N bond cleavage and double C(sp³)–C(sp³) bond formation within the iridium coordination sphere. This complex thus represents a unique functional analogue of methyltransferases and enables new synthetic transformations.

Previous studies in our laboratories have shown that [Ir(Cp*)Cl₂] reacts in the presence of Ag₂O with the pyridinium-triazolium salt 1 either via pyridinium C(sp³)–H bond activation or via exocyclic C–H bond activation to give 2 and 3, respectively (Scheme 2).[10] If acetate is added to the reaction mixture either as Ag₂OAc or NaOAc, 1 undergoes an N₆–C₉₅ bond activation process instead and affords complex 4a comprising a tridentate triazolylidene ligand with a chelating pyridine and imine donor group.[11] Formally, complex 4a is the product of a methylene shift from the pyridinium fragment to the benzyl group, followed by insertion of a MeCN molecule. Support for solvent activation[12] was obtained by carrying out the reaction in benzonitrile (PhCN) instead of MeCN, which yielded complex 4b.

In solution, complexes 4a and 4b each display two characteristic AB resonance patterns for the two pairs of benzylic protons (δ¹H 14.1 and 12.9 Hz, respectively, in 4a), which are in a rigidly fixed geometry due to the tridentate bonding of the ligand. The imine-bound proton appears at slightly lower field in 4a (δH 10.58) than in 4b (δH 9.85). In the ¹³C NMR spectrum, the N₆-donor-bound carbon is observed at 190 ppm. Most diagnostically, 1H spectroscopy revealed a stretch vibration at ν Chr ≈ 1635±1 cm⁻¹.

The connectivity pattern was confirmed by single crystal X-ray diffraction studies on 4a and 4b (Fig. 1). The NCN-tridentate ligand may be considered as facially coordinating pincer ligand.[13] The bond angles of the 5-membered metalacycle comprising the pyridine and the triazolylidene unit is acute and indicates some strain (Cpy–Ir–N₆ 75°), while the 9-membered metalacycle is considerably more flexible and adopts a coordination mode close to ideal for pseudo-octahedral iridium(III) complexes (Cpy–Ir–N₆ 85–90°). The higher strain is also reflected in the 0.07–0.09 Å

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Mechanistic details of this C–N bond breaking and multiple C–C bond making process were investigated by isotope labeling studies. When using compound I with a $^{13}$C-labeled methyl group at $N_{pyr}$ (15) under standard reaction conditions, complex 4a* was produced, which contained the $^{13}$C nucleus exclusively in the benzylic position between the aryl and the imine unit. Selective $^{13}$C labeling of 4a* was confirmed by the split of the resonance due to the two benzylic protons into two doublets of doublets ($\delta_1$ 4.27, $J_{CH} = 128$ Hz, $J_{HH} = 12.9$ Hz), and by the doublet resonance for the NH proton ($J_{CH} = 9$ Hz). Similarly, all $^{13}$C NMR signals due to the phenyl group and the C=N functionality appear as doublets as a consequence of their coupling to the benzylic $^{13}$C nucleus (e.g., $\delta_{C=N}$ 192.05 with $J_{CC} = 43.1$ Hz). No traces of unlabeled benzylic carbon were detected, indicating a selective transfer of the $N_{pyr}$-bound carbon to the phenyl ring of the benzylic substituent.

Deuterium labeling of 1 at both $N_{pyr}$ and $N_{ar}$ by using CD$_3$OTf as methylating agent gave 1–D$_3$. Reaction of this partially deuterated precursor with [Ir(Cp*)Cl]$_3$ in CH$_2$CN under standard conditions gave 4a with only little D-incorporation. The $^1$H NMR spectrum revealed 17% deuterium in the benzylic position as compared to the triazole-bound $N_{ar}$–CD$_3$ group, yet no deuterium incorporation in the terminal CH$_3$ group attached to the imine. The corresponding $^1$H NMR spectrum confirmed these measurements, indicating no $N_{ar}$–CH$_3$ residues and a benzylic proton integration of approximately 80%. The inverse labeling, i.e. starting from the proto ligand precursor 1 and performing the reaction in CD$_3$CN gave complex 4a–D$_3$ with essentially complete deuteriation at the terminal methyl group bound to the imine, and approximately 60(±10)% deuteration incorporation into the benzylic position. While these results do not allow for determining whether a methyl or a methylene group is transferred, they clearly indicate solvent-mediated isotope scrambling during the transfer process. No such scrambling at the methylene group interlinking the aryl and the triazole heterocycle or at the $N_{ar}$-bound methyl group was observed. Hence, isotope exchange in the starting material seems unlikely. Similarly, exchange processes after the formation of 4 at only one of the two available benzylic positions is not supported.

When the reaction was stopped before completion, a variety of intermediates were detected. Thus after 2 h, a hydride-containing intermediate was observed ($\delta_1$ –14.28 ppm, Cp* protons appear as doublet with $J_{HH} = 0.8$ Hz). Separation of the product mixture at this stage failed to give the hydride species in pure form. However, two species were isolated that were assigned to B and B' along with minor quantities of 4a (Scheme 3). These products are present in a 1:0.7 ratio, irrespective of the reaction time (2–16 h) or the reaction temperature (25–85 °C), pointing to a thermodynamically controlled distribution. The two species are similar according to their $^1$H NMR spectrum, both containing four pyridyl signals, two methyl groups (for $N_{pyr}$–CH$_3$ and $N_{ar}$–CH$_3$), and the benzylic protons split into an AB signal. Most diagnostically, both compounds contain only four phenyl protons appearing as two doublets and two doublets of doublets, which suggests orthometallation via C$_{benzylic}$–H bond activation of the benzylic group. The major difference between the two species consists in the chemical shift of the $N_{pyr}$–CH$_3$ group ($\delta_1$ 4.37 vs 3.92) and the meta-positioned pyridinium proton (C3–H at $\delta_1 8.08$ vs 8.33). These differences concur with the presence of two rotaers comprising the $N_{pyr}$–CH$_3$ group either pointing towards (B) or away from the monodentate NCMe ligand at iridium (B'). Nuclear Overhauser experiments indicate an anti conformation of the two N–CH$_3$ groups for the major isomer.

These cyclometalated products were also obtained from I and [Ir(Cp*Cl)$_2$]OTf anhydride only, i.e. in the absence of AgOAc. In CH$_2$CN, the stability of the intermediates B and B' is greatly enhanced as C–H bond activation and ylide formation is only induced after several days at reflux temperature, indicating that NCMe displacement from B is essential to form complex 3. The ylide complex 3 is stable when heated in MeCN in the presence of OAc$^–$. In contrast, the carbene intermediate B and complex 2 undergo methylene transfer under these conditions, thus gradually generating 4a.

A tentative mechanism that is compliant with these observations is depicted in Scheme 3. Intermediate A has been observed previously and may be formed via in-situ transmetalation or by iridium-mediated C–H activation, which would rationalize the traces of iridium-hydride species observed. Subsequent cyclometalation, probably OAc$^–$ assisted or via oxidative addition generates a mixture of the C,C-bidentate complex B and B'. Upon exchange of the MeCN ligand in B by acetate, concerted and supposedly rate-limiting activation of the C–H bond and cleavage of the $N_{pyr}$–C bond ensues (C), thus producing a carbene species (speculatively represented as D) which may be susceptible to protonation and solvent-mediated H/D exchange. The interplay of acetate and iridium in mediating the proton abstraction and $N_{pyr}$–C bond activation seems most critical to this methyl transfer process. Methylene insertion into the Ir–C$_{aryl}$ bond, followed by activation of a coordinated solvent molecule via nucleophilic addition of the anionic benzylic group is postulated to generate the nine-membered metallacycle in 4. Albeit tentative, this mechanism takes into consideration that acetate is essential for the reaction to occur, and it allows the formation of the ylide 3 from the intermediates B and B' to be rationalized in the absence of acetate. The transfer of the carbon is selective, while H/D scrambling with the solvent may occur either at the carbene intermediate D or before protonation of the imide ligand in the conversion of E to 4. Due to the fast proton exchange at acetate, the mechanism also provides a rationale for the fact that none of the deuterium labeling experiments resulted in D-incorporation at the imine position.

In agreement with the proposed model, the methylene transfer process is suppressed when the precursor I contains a fluorinated benzylic group (CH$_2$CF$_2$). No products similar to 4 were observed.
and instead, only pyridinium C–H bond activation took place to give C₃F₇-containing analogues of 2 and 3. Likewise, substitution of the benzylic group in 1 with a phenyl unit suppressed the methylene group transfer and afforded a bidentate cyclometalated product resulting from C₃phenyl-H bond activation reminiscent of B.[20]

Apparently, the steric flexibility of the benzyl group promotes the carbon transfer while phenyl coordination induces sufficient constraints to prevent the pyridinium ring from approaching the iridium center. Attempts to expand the reaction towards the transfer of different alkyl groups were unsuccessful. When using the ethyl-pyridinium analogue of 1, a complicated mixture of products formed that was inseparable in our hands, yet the crude product mixture showed no signals that might indicate the migration of the ethyl group from the pyridinium fragment.

Complex 4a is stable under normal conditions and undergoes only incomplete N–H to N–D exchange in the presence of D₂O even after several days. In acidic media (methanolic HCl), rapid nitrile activation and subsequent C(sp²)–H bond formation in the metal-catalyzed cross-coupling of unfunctionalized arenes.[12] and nitrile activation in recent metal-mediated reactions.[12] Combining these processes in a single transformation provides a first functional model of methyltransferase and opens new avenues for organic functionalizations.

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indicating a pseudo-N(Noe) upon saturation at the N
minor component through Saturation of the major N
and at 5.66 and 5.13 ppm for The AB sets are centered at 5.66 and 5.09 ppm for suggest this fragment to be involved in the alkyl transfer mechanism.

M. Albrecht, G. van Koten, Togni, Boyle, J. L. Petersen, 2002

MeCN activation has precedents, see for examples: a) S. D. Ittel, C. A. See the supporting information for details.


Variable temperature NMR spectroscopy between −30 and +70 °C did not provide any evidence for interconversion of the two species.


Hop on — hop off: an iridium center transfers a methyl group from pyridinium to an aryl unit, using exclusively the pyridine-bound methyl group as a mild methylene source and accomplishing the cleavage of an unactivated C(aryl)–H bond and nitrile solvent activation. The process is reminiscent of DNA methylation and entails the formation of two new C(sp²)–C(sp³) bonds within the metal coordination sphere.